



General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

Phase Changes

$$\Delta T_b = iK_b m$$

$$\Delta T_f = iK_f m$$

$$\pi = iMRT$$

$$c = kP$$

$$\frac{P_1}{P_2} = \frac{r_1}{r_2}$$

$$P_{\text{solution}} = \chi_{\text{solvent}} P_{\text{solvent}}^{\circ}$$

$$\Delta P = P_{\text{solvent}}^{\circ} - P_{\text{solution}} = \chi_{\text{solute}} P_{\text{solvent}}^{\circ}$$

Cells

$$E = E^{\circ} - \frac{RT}{nF} \ln(Q)$$

$$E = E^{\circ} - \frac{0.0592}{n} \log(Q) \quad (\text{at } 25^{\circ}\text{C})$$

$$\Delta G = -nFE_{\text{cell}}$$

$$\Delta G^{\circ} = -nFE_{\text{cell}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = \frac{RT}{nF} \ln(K)$$

$$E_{\text{cell}}^{\circ} = \frac{0.0592\text{V}}{n} \log(K) \quad (\text{at } 25^{\circ}\text{C})$$

USE ONE:

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re d.)}}^{\circ} - E_{\text{anode(re d.)}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re d.)}}^{\circ} + E_{\text{anode(ox.)}}^{\circ}$$

Thermodynamics

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta S_{\text{surr}} = \frac{-\Delta H_{\text{sys}}}{T}$$

$$\Delta S_{\text{univ}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$$

$$\Delta S_{\text{ran}}^{\circ} = \Sigma S_{\text{products}}^{\circ} - \Sigma S_{\text{reactants}}^{\circ}$$

$$\Delta G_{\text{ran}}^{\circ} = \Sigma \Delta G_{\text{products}}^{\circ} - \Sigma \Delta G_{\text{reactants}}^{\circ}$$

Concentrations

$$\text{molality} = \frac{\text{moles solute}}{\text{mass solvent (kg)}}$$

$$\% \text{ by mass} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 100\%$$

$$\text{parts per million} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 10^6$$

$$\text{parts per billion} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 10^9$$

Acids and Bases

$$K_w = [H_3O^+][OH^-] = K_a \times K_b$$

$$pOH = -\log[OH^-]$$

$$[OH^-] = 10^{-pOH}$$

$$pH + pOH = pK_a + pK_b = 14.00 \quad (\text{at } 25^{\circ}\text{C})$$

$$pH = pK_a + \log \frac{[A^-]}{[HA]}$$

$$pOH = pK_b + \log \frac{[BH^+]}{[B]}$$

$$\% \text{ ionization} = \frac{[H_3O^+]_{\text{eq}}}{[HA]_{\text{in}}} \times 100\%$$

$$pK_a = -\log(K_a)$$

$$pK_b = -\log(K_b)$$

$$K_a = 10^{-pK_a}$$

$$K_b = 10^{-pK_b}$$

Equilibria

$$Q_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

$$K_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

$$Q_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = K_c [R \times T]^{\Delta n}$$

$$\Delta G = \Delta G^{\circ} + RT \ln(Q)$$

$$\Delta G^{\circ} = -RT \ln(K)$$

$$\text{rate} = k[A]^x[B]^y$$

$$k = Ae^{-\frac{E_a}{RT}}$$

$$\ln\left(\frac{k_2}{k_1}\right) = \frac{E_a}{R} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)$$

$$\text{half life} = \frac{\ln(2)}{k} \quad (\text{1st order reactions only})$$

$$[A]_t = -kt + [A]_0$$

$$\ln([A]_t) = -kt + \ln([A]_0)$$

$$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$$

Answer Sheet General Chemistry Equations

Leo J. Malone, Theodore Dolter



Answer Sheet General Chemistry Equations:

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Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value

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