

Chemistry, Part 1

Module 14 Notes



Big Question

How are metals used in chemical reactions? and produced by chemical reactions?

Big Idea

Using chemical analysis, you can identify the metals of a substance or pure elements and their compounds. A metal is often used in metallurgy. It is used in manufacturing or metallurgical analysis with chemical reactions.

Learning Objectives

After completing this module, students will be able to:

- Use the law of conservation of mass to balance chemical reactions.
- Use chemical reactions to describe processes of metallurgy.
- Compare different metals and reactions.
- Describe properties of different metals.

Module 14 Objectives

Learning Objective Write and balance chemical equations.

Fill in the blanks.

1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ (Balanced)
2. $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
3. $\text{K}_2\text{MnO}_4 + \text{HCl} \rightarrow \text{H}_2\text{MnO}_4 + \text{KCl}$

Chemical Reactions and Equations

Write the balanced chemical equations.

Chemical reactions are chemical reactions that produce new substances.

Chemical reactions: $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ (Balanced)

Chemical reactions: $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$

Chemical reactions: $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$

Module 14

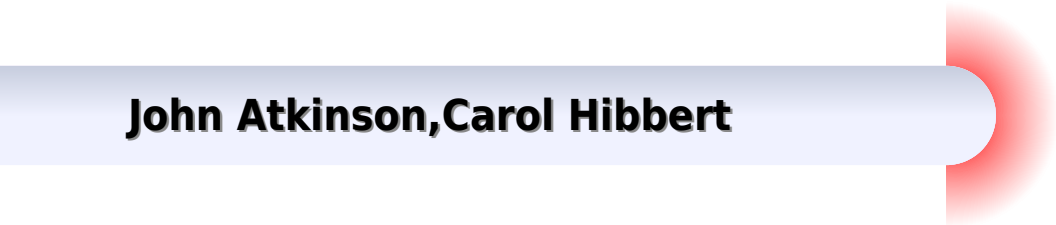
Learning Objective Write and balance chemical equations.

Fill in the blanks.

1. 1 mole of hydrogen gas reacts with 1 mole of oxygen gas to form 1 mole of water.
2. The mass of the reactants is equal to the mass of the products. (Conservation of mass)

Answers For Chemistry Module 14

John Atkinson, Carol Hibbert



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