



basic education

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GRADE 12

GEOGRAPHY P1

MEMORANDUM

FEBRUARY/MARCH 2014

MARKS: 300

This memorandum consists of 15 pages.

2014 Grade 12 Preliminary Memo Tshwane District 2014

Daniela Niemeyer



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Table of Contents 2014 Grade 12 Preliminary Memo Tshwane District 2014

1. Understanding the eBook 2014 Grade 12 Preliminary Memo Tshwane District 2014
 - The Rise of Digital Reading 2014 Grade 12 Preliminary Memo Tshwane District 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying 2014 Grade 12 Preliminary Memo Tshwane District 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2014 Grade 12 Preliminary Memo Tshwane District 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2014 Grade 12 Preliminary Memo Tshwane District 2014
 - Personalized Recommendations
 - 2014 Grade 12 Preliminary Memo Tshwane District 2014 User Reviews and Ratings
 - 2014 Grade 12 Preliminary Memo Tshwane District 2014 and Bestseller Lists
5. Accessing 2014 Grade 12 Preliminary Memo Tshwane District 2014 Free and Paid eBooks
 - 2014 Grade 12 Preliminary Memo Tshwane District 2014 Public Domain eBooks
 - 2014 Grade 12 Preliminary Memo Tshwane District 2014 eBook Subscription Services
 - 2014 Grade 12 Preliminary Memo Tshwane District 2014 Budget-Friendly Options
6. Navigating 2014 Grade 12 Preliminary Memo Tshwane District 2014 eBook Formats

- ePub, PDF, MOBI, and More
 - 2014 Grade 12 Preliminary Memo Tshwane District 2014 Compatibility with Devices
 - 2014 Grade 12 Preliminary Memo Tshwane District 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 2014 Grade 12 Preliminary Memo Tshwane District 2014
 - Highlighting and Note-Taking 2014 Grade 12 Preliminary Memo Tshwane District 2014
 - Interactive Elements 2014 Grade 12 Preliminary Memo Tshwane District 2014
8. Staying Engaged with 2014 Grade 12 Preliminary Memo Tshwane District 2014
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2014 Grade 12 Preliminary Memo Tshwane District 2014
9. Balancing eBooks and Physical Books 2014 Grade 12 Preliminary Memo Tshwane District 2014
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2014 Grade 12 Preliminary Memo Tshwane District 2014
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2014 Grade 12 Preliminary Memo Tshwane District 2014
- Setting Reading Goals 2014 Grade 12 Preliminary Memo Tshwane District 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2014 Grade 12 Preliminary Memo Tshwane District 2014
- Fact-Checking eBook Content of 2014 Grade 12 Preliminary Memo Tshwane District 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno₃ h₂so₄ in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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