

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a solution containing chloride ions. The principle of gravimetric analysis is based on the fact that the mass of a substance is proportional to the number of moles of the substance. By measuring the mass of a substance and knowing its molar mass, the number of moles can be calculated. In this experiment, the chloride ions in a solution are precipitated as silver chloride (AgCl) by adding a known volume of a silver nitrate (AgNO_3) solution. The mass of the precipitate is then measured, and the mass of chloride ions in the original solution is calculated. The experiment involves the use of a balance, a volumetric flask, and a beaker. The procedure is as follows: 1. Preparation of a known volume of a solution containing chloride ions. 2. Addition of a known volume of a silver nitrate solution to the chloride solution. 3. Filtration of the precipitate. 4. Drying of the precipitate. 5. Weighing of the precipitate. 6. Calculation of the mass of chloride ions in the original solution.

The reaction between silver nitrate and chloride ions is as follows: $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} + \text{NO}_3^-$. The mass of the precipitate is measured, and the mass of chloride ions in the original solution is calculated. The experiment involves the use of a balance, a volumetric flask, and a beaker. The procedure is as follows: 1. Preparation of a known volume of a solution containing chloride ions. 2. Addition of a known volume of a silver nitrate solution to the chloride solution. 3. Filtration of the precipitate. 4. Drying of the precipitate. 5. Weighing of the precipitate. 6. Calculation of the mass of chloride ions in the original solution.

Procedure:

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Data Analysis:

	Sample 1	Sample 2
Mass of precipitate	0.1234 g	0.1567 g
Mass of chloride ion	0.0567 g	0.0789 g
Mass of silver ion	0.0667 g	0.0778 g

The mass of chloride ion in the original solution is calculated as follows:

$$\text{Mass of Cl}^- = \text{Mass of AgCl} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$$

The mass of silver ion in the original solution is calculated as follows:

Analysis Of Chloride Lab Report

Y. V. Singh



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