

**Analytical Chemistry**  
**3<sup>rd</sup> Long Exam**

**PART I. MULTIPLE CHOICE.** Choose the letter that corresponds to the **BEST** answer. Write **[ ]** when the answer is **NOT** in the choices. (25 pts)

1. Transfer of electrons from one reactant to another

- a. precipitation rxn                      c. redox rxn  
b. half rxn                                  d. neutralization rxn

2. What is the mole ratio of  $\text{Ce}(\text{IV})$  to  $\text{NO}_2^-$  for the oxidation of nitrite to nitrate using  $\text{Ce}^{4+}$  in the presence of an acid:



- a. 1:2                                          c. 1:1  
b. 2:1                                          d. 2:2

3. How many mole(s) of electron is(are) produced in the oxidation of sodium bisulfate ( $\text{NaHSO}_4$ ) to  $\text{BiO}^+$  and  $\text{H}^+$  under acidic conditions?

- a. 1                                              c. 3  
b. 2                                              d. 4

4. Azide ( $\text{N}_3^-$ ) was added in the determination of the dissolved oxygen in water sample to remove interferences caused by the presence of what ion?

- a. chloride                                  c. nitrite  
b. fluoride                                  d. nitrate

5. Which of the following statements is(are) true?

- I. Prepared  $\text{KMnO}_4$  solutions should be stored in amber bottles  
II.  $\text{KMnO}_4$  solutions are boiled to remove  $\text{CO}_2$  which makes the solution basic  
III.  $\text{Mn}(\text{II})$  is the product of the degradation of  $\text{KMnO}_4$   
IV.  $\text{KMnO}_4$  solutions are filtered to remove manganese(IV) oxide

- a. I only                                      c. I and IV  
b. I and II                                  d. I, II and III

6. What is the oxidation state of carbon in oxalic acid?

- a. +2                                          c. +6  
b. +4                                          d. +7

7. The complex formed between  $\text{Cu}(\text{I})$  and 1,10-phenanthroline has a molar absorptivity of  $7000 \text{ L cm}^{-1} \text{ mol}^{-1}$  at 435 nm, the wavelength of maximum absorption. What is the absorbance of a  $5.17 \times 10^{-3} \text{ M}$  solution of the complex when measured in a 1.00-cm cell at 435 nm?

- a. 0.0431                                      c. 0.432  
b. 0.0432                                      d. 0.431

8. In the presence of iodine, thiosulfate ion is oxidized quantitatively to tetrathionate ion. What is the chemical formula of tetrathionate ion?

- a.  $\text{S}_4\text{O}_7^{2-}$                                       c.  $\text{S}_4\text{O}_6^{2-}$   
b.  $\text{S}_4\text{O}_5^{2-}$                                       d.  $\text{S}_4\text{O}_8^{2-}$

9. In Winkler titration, the flocc is produced from

- a. oxidation of  $\text{I}^-$  to  $\text{I}_2$                       c. reduction of  $\text{I}^-$  to  $\text{I}_2$   
b. oxidation of  $\text{Mn}(\text{II})$  to  $\text{Mn}(\text{III})$               d. reduction of  $\text{Mn}(\text{II})$  to  $\text{Mn}(\text{III})$

10. A component of a UV-Vis spectrophotometer that restricts the radiation being measured to a narrow band that is absorbed or emitted by the analyte.

- a. light source                                  c. radiation filter  
b. detector                                      d. wavelength selector

11. Component of a UV-Vis instrument that operates by blocking or absorbing all but a restricted band of radiation.

- a. light source                                  c. radiation filter  
b. detector                                      d. wavelength selector

12. Oxidation of  $\text{VO}^{2+}$  using permanganate or cerium as standard solution produces  $\text{VO}(\text{OH})_2^{+}$ , what is the oxidation state of vanadium in  $\text{VO}(\text{OH})_2^{+}$ ?

# Analytical Chemistry Exam Questions

**Arshad Iqbal**



## **Analytical Chemistry Exam Questions:**

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*Understanding Advanced Organic And Analytical Chemistry: The Learner's Approach (Revised Edition)* Kim Seng Chan, Jeanne Tan, 2016-09-29 This revised edition has been updated to meet the minimum requirements of the new Singapore GCE A level syllabus that would be implemented in the year 2016 Nevertheless this book is also highly relevant to students who are studying chemistry for other examination boards In addition the authors have also included more Q A to help students better understand and appreciate the chemical concepts that they are mastering

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