



ANALYZING THE “POP” IN POPCORN

LABORATORY EXPERIMENT

ADAPTED FROM THE ACS CHEMSOURCE GAS MODULE

INTRODUCTION

Corn is a common foodstuff, native to the Americas, which appears in many forms: corn on the cob, corn off the cob, creamed corn, and popcorn. Field corn, which is also fed to livestock, is used by some people as decorations during Halloween and Thanksgiving. Popcorn is a favorite snack, especially when watching movies. Each variety of corn contains different amounts of water, sugars, and starches. Popping popcorn involves heating the corn until the pressure inside the kernel is great enough to cause it to burst, turning the kernel inside out and releasing trapped moisture.

PURPOSE

To determine the percent of water in popcorn and use the ideal gas law to determine the pressure inside the kernel when it pops.

SAFETY INFORMATION

In this lab, observe all standard safety precautions, especially the ones listed below.

- Safety goggles and lab aprons must be worn at all times.
- Heat the flask evenly to prevent spattering the oil.
- Heat with care. Oil is flammable.
- Do not eat the popcorn.

PROCEDURES

1. Record the mass of a plastic weighboat.
2. Select and weigh 20 kernels of one type of popcorn.
3. Calculate the mass of the popcorn kernels.
4. Using the water displacement method, find the volume of the 20 kernels using a 10-mL graduated cylinder filled with 5-mL of water. Dry the kernels thoroughly.
5. Add 2 mL of cooking oil and the 20 kernels to an empty, dry 250-mL Erlenmeyer flask.
6. Determine the mass of the flask, oil, and popcorn.
7. Assemble a ringstand with ring, placing wire gauze on top of the ring.
8. Using a utility clamp, fasten the flask to the stand so it sits on the wire gauze. Cover the flask with a piece of aluminum foil. Puncture a small hole in the aluminum foil.
9. Light the burner.
10. Use the burner to **slowly** heat the flask until the popcorn begins to pop. Remove the heat when most of the kernels have popped. Do not burn the popcorn.
11. Remove the aluminum foil from the top of the flask so all the steam vents out.
12. Allow the flask to cool for 10 minutes, and then weigh the flask and contents.
13. Clean up the lab as follows:
 - a. Place the aluminum foil and popcorn in the garbage.
 - b. Wash the flask and graduated cylinder thoroughly with soap and water.
 - c. Do not leave your lab bench or remove your safety clothing until excused by the instructor.

Analyzing The Pop In Popcorn Lab Answers

Michal Rosen-Zvi



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Resources in Education, 1975 *Billboard*, 1945-09-01 In its 114th year Billboard remains the world's premier weekly music publication and a diverse digital events brand content and data licensing platform Billboard publishes the most trusted charts and offers unrivaled reporting about the latest music video gaming media digital and mobile entertainment issues and trends
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Report of the Division of Chemistry Canada. Department of Agriculture. Division of Chemistry, 1924
American Book Publishing Record, 1977-03-31 Here's quick access to more than 490 000 titles published from 1970 to 1984 arranged in Dewey sequence with sections for Adult and Juvenile Fiction Author and Title indexes are included and a Subject Guide correlates primary subjects with Dewey and LC classification numbers These cumulative records are available in three separate sets
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Progress Report for the Years Canada. Department of Agriculture. Division of Chemistry, 1912
Subject Guide to Books in Print, 1983
Popular Science, 1947-11 Popular Science gives our readers the information and tools to improve their technology and their world The core belief that Popular Science and our readers share The future is going to be better and science and technology are the driving forces that will help make it better
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Snack Food, 1989
Flying Magazine, 1966-11
Billboard Music Week, 1941
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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $\frac{160 \text{ g Al}}{27 \text{ g/mol Al}} \times \frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.96 \text{ g/mol Al}_2\text{O}_3 = 302.3 \text{ g}$ c determine the percent yield $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260 \text{ g}}{302.3 \text{ g}} \times 100 = 86.0\%$

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