

1. Answer: A

**Explanation:** The value of  $\Delta H$  is negative. Therefore, the reaction is exothermic or heat-releasing.

2. Answer: A

**Explanation:** The entire reaction was multiplied by  $\frac{1}{2}$ . Therefore,  $\Delta H$  must also be halved.

3. Answer: C

**Explanation:** Option C is incorrect because  $-566 \text{ kJ/mol}$  is the  $\Delta H$  per mole of  $O_2$  and not for 2 moles. Even though options B and D are expressed in grams, they are still correct because 56 g of CO still corresponds to 2 moles of CO, and 32 g of  $O_2$  still corresponds to 1 mole of  $O_2$ , which is consistent with the expected factor.

4. Answer: A

**Explanation:**  $\Delta H = (98 \text{ g } CO_{(g)}) \left( \frac{1 \text{ mol } CO_{(g)}}{28 \text{ g } CO_{(g)}} \right) \left( \frac{-566 \text{ kJ}}{1 \text{ mol } CO_{(g)}} \right) = -990.5 \text{ kJ}$

5. Answer: B

**Explanation:**  $\Delta H = (1.5 \text{ mol } CO_{2(g)}) \left( \frac{-566 \text{ kJ}}{1 \text{ mol } CO_{2(g)}} \right) = -424.5 \text{ kJ}$

# Answer Key Unit 6 Thermochemistry

**National Institute of Standards and  
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