

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Sub-levels have different amounts of energy. For example, orbitals closer to the nucleus have more energy than orbitals further from the nucleus. The following orbitals have different levels of energy:



For each pair, write sublevel(s) for which orbitals are included. Then draw orbital(s) representing an orbital.



Now list 4 orbitals which appear in the set of pairs. Write down their approximate energy levels. For example, the 3d orbitals have higher energy than s orbitals and the 4f orbitals have a higher energy than d orbitals.

When electrons occupy orbitals, they try to have the lowest amount of energy possible. (This is called the **lowest-energy principle**.) An electron will occupy 1 orbital before the 1 orbital is filled and no electron will occupy an orbital until it is completely filled. Additionally, no electron will be sub-filled. (This is called the **Mutual Exclusion Principle**.) When two electrons occupy the same orbital they must spin in opposite directions – one clockwise and the other anticlockwise.

Critical Thinking Questions

- a) How many electrons would it take to fill a sub-level before beginning to be sub-filled? **Sub-levels become filled if the orbitals are filled before electrons reach orbitals in the 2p orbitals.**
b) What is the first element before enough electrons are present to fill 1 orbital? **After you provide this 1.**
The 1st 2p electron goes into the 1 orbital. The element is 1, hydrogen 1 (H).
- a) How many electrons would it take to fill a sub-level before beginning to be sub-filled? **Sub-levels become filled if the orbitals are filled before electrons reach orbitals in the 2p orbitals.**
b) What is the first element before enough electrons are beginning to fill orbitals in the 2p orbitals? **The 2nd 2p electron goes into the 2 orbitals. The element is 2, helium 2 (He).**

Information: Hund's Rule

Electrons prefer to "spin" or "upspin". Orbitals closer to the nucleus have a higher potential. Electrons closer to the nucleus have a lower potential. Electrons closer to the nucleus have a higher potential.



If an orbital is represented as a box, the orbital diagram will get electron's orbital diagram. The diagram represents a box for each orbital. The diagram represents a box for each orbital.



Hund's rule tells us that orbitals closer to the nucleus have a higher potential. Electrons closer to the nucleus have a higher potential. Electrons closer to the nucleus have a higher potential. Electrons closer to the nucleus have a higher potential.

The following are the electron orbital diagrams for the second element, nitrogen and oxygen. Write the diagram's 2p electrons and explain to show Hund's rule. The 2p electrons are located in the 2p orbitals.



Critical Thinking Questions

- How many "upspin" electrons are in a nitrogen atom? **1**
- Why does carbon's orbitals diagram is given as a 2p orbital? Why not a 2p orbital? **Why not a 2p orbital? The 2p orbitals are already full. The 2p orbitals are already full. The 2p orbitals are already full.**
- Why do electrons fill orbitals for nitrogen? **1**
- Why do electrons fill orbitals for oxygen? **1**

Answers For Chemquest

S Baum



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