

CHAPTER 10: Reactions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

1. For elements in their standard state, the value of ΔH_f° is _____.
2. The standard heat of decomposition of a substance is determined by the following thermodynamic equation:



- a. Describe how the value of $\Delta H_{\text{decomp}}^\circ$ is being interpreted in the above equation.
- b. What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

3. $\Delta H_{\text{decomp}}^\circ$ for a substance is 2074°C . Assume a 1.000 g of substance is heated until the decomposition has:

- a. $\Delta H_{\text{decomp}}^\circ$ is exactly half of 2074°C . Write the correct value for $\Delta H_{\text{decomp}}^\circ$ in standard scientific notation.
- b. $\Delta H_{\text{decomp}}^\circ$ is exactly half of 2074°C .

4. The product is a molecular form of oxygen, and the element has an enthalpy of 0 kJ/mol.

- a. What is the value of $\Delta H_{\text{decomp}}^\circ$ for this reaction?

Answers For Reaction Energy Section 1

Robert H. Crabtree



Answers For Reaction Energy Section 1:

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