

SCH4U Unit 4 - Quiz – Equilibrium

Name: _____ Period: _____

1. (20)

2. (20)

3. (20)

4. (20)

5. (20)

6. (20) (20) (20)

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{(P)^{\text{products}}}{(P)^{\text{reactants}}}$$

$$K_c = \frac{P_{\text{H}_2\text{O}}^2 P_{\text{CO}_2}}{P_{\text{CH}_4} P_{\text{CO}}^2} \quad K_p = \frac{(P_{\text{H}_2\text{O}})^2 (P_{\text{CO}_2})}{(P_{\text{CH}_4}) (P_{\text{CO}})^2}$$

gases: $\Delta G^\circ = -RT \ln K$ $\Delta G^\circ = -RT \ln K_p$ $\Delta G^\circ = -RT \ln K_c$ $\Delta G^\circ = -RT \ln K$ $\Delta G^\circ = -RT \ln K$

Question 1

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000°C, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.020 atm. The reaction is:



If the volume of an 8.00 L container at 1000°C where a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
- What is the concentration of S_2 at equilibrium?
- What is the concentration of H_2S at equilibrium?
- What is the value of K_c at equilibrium?
- What other species have chemical formula consisting of 2 atoms each?
- Calculate:

	$\text{H}_2(\text{g})$	$\text{S}_2(\text{g})$	$\text{H}_2\text{S}(\text{g})$
Initial (moles)			
Change (moles)			
Final (moles)	0.020	0.010	0.020

7. (20) Equilibrium constants:

Answers For Sch4u Unit 4

Michal Rosen-Zvi



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