

Name _____ answer key _____ Date _____

Stoichiometry: Mass – Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₃.

$$\frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3} \times \frac{1 \text{ mole FeCl}_3}{9 \text{ FeCl}_3} \times \frac{\text{mole NaCl}}{\text{mole FeCl}_3}$$

Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

$$\frac{9 \text{ Na}_2\text{SO}_4}{1 \text{ mole Na}_2\text{SO}_4} \times \frac{1 \text{ mole Na}_2\text{SO}_4}{9 \text{ Na}_2\text{SO}_4} \times \frac{\text{mole Fe}_2(\text{SO}_4)_3}{\text{mole Na}_2\text{SO}_4} \times \frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3}$$

Calculate the amount in grams of FeCl₃ produced with 95g of Fe (SO₄)₃

$$\frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3} \times \frac{1 \text{ mole Fe}_2(\text{SO}_4)_3}{9 \text{ Fe}_2(\text{SO}_4)_3} \times \frac{\text{mole FeCl}_3}{\text{mole Fe}(\text{SO}_4)_3} \times \frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3}$$

Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl

$$\frac{9 \text{ NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole NaCl}}{9 \text{ NaCl}} \times \frac{\text{mole Fe}(\text{SO}_4)_3}{\text{mole NaCl}}$$

Based on the following chemical equation, please answer the following questions!



$$\frac{9 \text{ Ca}(\text{HCO}_3)_2}{1 \text{ mole Ca}(\text{HCO}_3)_2} \times \frac{1 \text{ mole Ca}(\text{HCO}_3)_2}{9 \text{ Ca}(\text{HCO}_3)_2} \times \frac{\text{mole CaCO}_3}{\text{mole Ca}(\text{HCO}_3)_2}$$

Answers For Study Guide Mass And Moles

H. Stephen Stoker, Danny V. White



Answers For Study Guide Mass And Moles:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

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Frentrup, 2002-06 Contains a brief overview of every chapter review of skills, self-tests, and the answers and detailed solutions to all odd-numbered end-of-chapter problems in the textbook. **Chemistry, Student Study Guide** James E. Brady, Fred

Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane, CH_4 , from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. "O" Level Study Guide

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