

Guided Reading Activity 9-1

For use with textbook pages 211–217.

Population Patterns

Fill in the Blanks

DIRECTIONS: Use the information in your textbook to fill in the blanks for the following sentences.

1. Three Native American groups to develop civilizations in Latin America were the Maya of Central America, the Aztec of Mexico, and the Inca of Peru.
2. Of the first Europeans to settle in Latin America, most were Spanish and Portuguese.
3. Africans were brought to Brazil and the Caribbean to work on plantations.
4. Overall, the nations of Latin America have a diversity of ethnic groups.
5. Most Latin American countries adopted the language of their European colonizers.
6. People in different parts of Latin America speak different dialects, or forms of a language unique to a particular place or group.
7. Most Latin Americans live along the coasts of either the Atlantic or the Pacific Oceans.
8. The vast interior of South America remains largely uninhabited because of the harsh climate and rugged landscape.
9. The Brazilian government moved its capital from coastal Rio de Janeiro to Brasilia to encourage people to settle the interior.
10. Unlike South America, with its huge expanses of land, small Caribbean countries have extremely high population density.
11. A significant part of the Mexican population is concentrated in and around Mexico City.
12. A megacity, such as Mexico City, has expanded to absorb surrounding cities and suburb.
13. Primate cities are so big that they often dominate the country's politics, culture, and economy.
14. In Latin America, many poor people who flock to cities are forced to live in areas with substandard housing and few services.
15. The stress of poverty in overcrowded cities often results in homelessness for many children, who must fend for themselves on the streets.

Answers American History Guided Activity 16

Lawrence B. Glickman



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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
 8.3 mole H₂S x 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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 what is the limiting reagent if 8.52 g are actually formed what is the

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 step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.6656 \text{ moles O}_2$

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 yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction Al + O₂ → Al₂O₃
 4 Al + 3 O₂ → 2 Al₂O₃
 b determine the theoretical yield of Al₂O₃
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g/mol}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 150.1 \text{ g Al}_2\text{O}_3$
 c determine the percent yield
 $\frac{101.96 \text{ g}}{150.1 \text{ g}} \times 100 = 67.9\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield = 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H₂ are left over 13

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