



Exam-style questions and sample answers have been written by the authors. In examinations, the way marks are awarded may be different.

> Coursebook answers

Chapter 28

Science in context guidance

- Students should already be familiar with a number of wave behaviours such as reflection, refraction and diffraction. These are all phenomena that occur with light, but also with other waves such as water waves.
- Waves can also be transmitted and absorbed. It is worth asking students if they can explain wave phenomena using particles instead of waves. They may come up with some interesting ideas!

Self-assessment questions

1 energy $E = hf = 6.63 \times 10^{-34} \times 1.0 \times 10^{15}$
 $= 6.63 \times 10^{-19} \text{ J} = 6.6 \times 10^{-19} \text{ J}$

2 energy $E = \frac{hc}{\lambda}$
 For red light, $E = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{700 \times 10^{-9}}$
 $= 2.8 \times 10^{-19} \text{ J}$

For violet light, $E = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{400 \times 10^{-9}}$
 $= 5.0 \times 10^{-19} \text{ J}$

3 Rearrange energy $E = \frac{hc}{\lambda}$ to give
 wavelength $\lambda = \frac{hc}{E}$

a $\lambda = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{10^{-12}} = 2.0 \times 10^{-13} \text{ m}$,
 γ-ray

b $\lambda = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{10^{-15}} = 2.0 \times 10^{-16} \text{ m}$,
 X-ray

c $\lambda = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{10^{-18}} = 2.0 \times 10^{-7} \text{ m}$,
 ultraviolet

d $\lambda = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{10^{-20}} = 2.0 \times 10^{-4} \text{ m}$,
 infrared

e $\lambda = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{10^{-23}} = 2.0 \text{ m}$, radio
 wave

4 energy of one photon, $E = \frac{hc}{\lambda}$
 $= \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{6.48 \times 10^{-7}} = 3.07 \times 10^{-19} \text{ J}$

number of photons per second =
 $\frac{\text{power of laser}}{\text{energy per photon}} = \frac{0.0010}{3.07 \times 10^{-19}} = 3.26 \times$
 $10^{15} \text{ s}^{-1} \approx 3.3 \times 10^{15} \text{ s}^{-1}$

5 energy change, $W = 1.2 \text{ eV} = QV$
 $= 1.6 \times 10^{-19} \times 1.2 = 1.92 \times 10^{-19} \text{ J} \approx 1.9 \times 10^{-19} \text{ J}$
 energy in eV = 1.2 eV

6 energy in joules, $E = hf$
 $1 \text{ eV} = 1.60 \times 10^{-19} \text{ J}$
 $E = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^{15}}{1.60 \times 10^{-19}} = 12\,400 \text{ eV} \approx$
 12 keV

7 Rearrange energy $E = \frac{hc}{\lambda}$ to give wavelength λ
 $= \frac{hc}{E} = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^8}{10 \times 1.6 \times 10^{-19}} \approx 1.24 \times 10^{-7} \text{ m}$,
 ultraviolet

8 a energy gained = work done on proton by
 potential difference
 energy gained = $QV = 1.60 \times 10^{-19} \times 1500$
 $= 2.4 \times 10^{-16} \text{ J}$

b kinetic energy $E_k = \frac{1}{2}mv^2$ so speed $v =$
 $\sqrt{\frac{2E_k}{m}} = \sqrt{\frac{2 \times 2.4 \times 10^{-16}}{1.67 \times 10^{-27}}} =$
 $\approx 5.3 \times 10^5 \text{ m s}^{-1}$

9 Rearrange $eV = \frac{hc}{\lambda}$ to give Planck's constant,
 $h = \frac{eV\lambda}{c}$
 For infrared LED, $\lambda =$
 $\frac{1.6 \times 10^{-19} \times 1.35 \times 910 \times 10^{-9}}{3.0 \times 10^8} = 6.55 \times 10^{-34} \text{ J s}$



Answers For Learntci Chapter 28

Jacqueline Rhoades



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Table of Contents Answers For Learntci Chapter 28

1. Understanding the eBook Answers For Learntci Chapter 28
 - The Rise of Digital Reading Answers For Learntci Chapter 28
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Learntci Chapter 28
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Learntci Chapter 28
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Learntci Chapter 28
 - Personalized Recommendations
 - Answers For Learntci Chapter 28 User Reviews and Ratings
 - Answers For Learntci Chapter 28 and Bestseller Lists

5. Accessing Answers For Learntci Chapter 28 Free and Paid eBooks
 - Answers For Learntci Chapter 28 Public Domain eBooks
 - Answers For Learntci Chapter 28 eBook Subscription Services
 - Answers For Learntci Chapter 28 Budget-Friendly Options
6. Navigating Answers For Learntci Chapter 28 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Learntci Chapter 28 Compatibility with Devices
 - Answers For Learntci Chapter 28 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Learntci Chapter 28
 - Highlighting and Note-Taking Answers For Learntci Chapter 28
 - Interactive Elements Answers For Learntci Chapter 28
8. Staying Engaged with Answers For Learntci Chapter 28
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Learntci Chapter 28
9. Balancing eBooks and Physical Books Answers For Learntci Chapter 28
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Learntci Chapter 28
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Learntci Chapter 28
 - Setting Reading Goals Answers For Learntci Chapter 28
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Learntci Chapter 28
 - Fact-Checking eBook Content of Answers For Learntci Chapter 28
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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