

2. Metallic magnesium reacts with steam to produce magnesium hydroxide and hydrogen gas.

a. Write a balanced equation for this reaction. (Note: It is very important that you do this step correctly!)



b. What type of reaction is this? single replacement

c. If 16.2 g of Mg are heated with 12.0 g of H_2O , which is the limiting reactant?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.0 \text{ g H}_2\text{O}$$

There is less H_2O (12.0 g) in our initial amount so
 H_2O is the L.R

d. How many moles of the excess reactant remain after the reaction is complete?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} = 1.3 \text{ mol H}_2\text{O}$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} = 0.666 \text{ mol H}_2\text{O}$$

$$1.3 \text{ mol H}_2\text{O} - 0.666 \text{ mol H}_2\text{O} = 0.634 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Mg}}{2 \text{ mol H}_2\text{O}} = \boxed{0.317 \text{ mol Mg}}$$

e. How many grams of each product are formed?

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol Mg}(\text{OH})_2}{2 \text{ mol H}_2\text{O}} \times \frac{58.32 \text{ g Mg}(\text{OH})_2}{1 \text{ mol Mg}(\text{OH})_2} = 19.4 \text{ g Mg}(\text{OH})_2$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} \times \frac{2.02 \text{ g H}_2}{1 \text{ mol H}_2} = 6.28 \text{ g H}_2$$

e. You perform this reaction, and obtain 16.0 g $\text{Mg}(\text{OH})_2$. What is your percent yield?

$$\% \text{ yield} = \frac{\text{actual} \times 100}{\text{theoretical}}$$

$$\frac{16.0 \text{ g Mg}(\text{OH})_2}{19.4 \text{ g Mg}(\text{OH})_2} \times 100 = 82.5\%$$

Answer Sheet To Chemistry 1 Quiz Review Stoichiometry

Neil D. Jespersen, Pamela Kerrigan



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Laboratory Questions Chapter 10 Liquids and Solids Questions Chapter 11 Mole Concept Questions Chapter 12 Oxidation Reduction Questions Chapter 13 Rates of Reactions Questions Chapter 14 Solutions Questions Chapter 15 Thermochemistry Questions The Molecular Structure Quiz Questions PDF e Book Chapter 1 interview questions and answers on polarity three dimensional molecular shapes The Acids and Bases Quiz Questions PDF e Book Chapter 2 interview questions and answers on Arrhenius concept Bronsted lowry concept indicators introduction Lewis concept pH strong and weak acids and bases The Atomic Structure Quiz Questions PDF e Book Chapter 3 interview questions and answers on electron configurations experimental evidence of atomic structure periodic trends quantum numbers and energy levels The Bonding Quiz Questions PDF e Book Chapter 4 interview questions and answers on ionic bond covalent bond dipole dipole forces hydrogen bonding intermolecular forces London dispersion forces metallic bond The Chemical Equations Quiz Questions PDF e Book Chapter 5 interview questions and answers on balancing of equations limiting reactants percent yield The Descriptive Chemistry Quiz Questions PDF e Book Chapter 6 interview questions and answers on common elements compounds of environmental concern nomenclature of compounds nomenclature of ions organic compounds periodic trends in properties of the elements reactivity of elements The Equilibrium Systems Quiz Questions PDF e Book Chapter 7 interview questions and answers on equilibrium constants introduction Le chatelier s principle The Gases Quiz Questions PDF e Book Chapter 8 interview questions and answers on density gas law relationships kinetic molecular theory molar volume stoichiometry The Laboratory Quiz Questions PDF e Book Chapter 9 interview questions and answers on safety analysis experimental techniques laboratory experiments measurements measurements and calculations observations The Liquids and Solids Quiz Questions PDF e Book Chapter 10 interview questions and answers on intermolecular forces in liquids and solids phase changes The Mole Concept Quiz Questions PDF e Book Chapter 11 interview questions and answers on Avogadro s number empirical formula introduction molar mass molecular formula The Oxidation Reduction Quiz Questions PDF e Book Chapter 12 interview questions and answers on combustion introduction oxidation numbers oxidation reduction reactions use of activity series The Rates of Reactions Quiz Questions PDF e Book Chapter 13 interview questions and answers on energy of activation catalysis factors affecting reaction rates finding the order of reaction introduction The Solutions Quiz Questions PDF e Book Chapter 14 interview questions and answers on factors affecting solubility colligative properties introduction molality molarity percent by mass concentrations The Thermochemistry Quiz Questions PDF e Book Chapter 15 interview questions and answers on heating curves calorimetry conservation of energy cooling curves enthalpy heat changes enthalpy heat changes associated with phase changes entropy introduction specific heats [Class 11-12 Chemistry MCQ \(Multiple Choice Questions\)](#) Arshad Iqbal,2019-05-17 The Class 11 12 Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF College Chemistry MCQ PDF Download Quiz Questions Chapter 1 6 Practice Tests with Answer Key 11th 12th Grade Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs Class 11

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Table of Contents Answer Sheet To Chemistry 1 Quiz Review Stoichiometry

1. Understanding the eBook Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
 - The Rise of Digital Reading Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
 - Personalized Recommendations
 - Answer Sheet To Chemistry 1 Quiz Review Stoichiometry User Reviews and Ratings
 - Answer Sheet To Chemistry 1 Quiz Review Stoichiometry and Bestseller Lists
5. Accessing Answer Sheet To Chemistry 1 Quiz Review Stoichiometry Free and Paid eBooks

- Answer Sheet To Chemistry 1 Quiz Review Stoichiometry Public Domain eBooks
 - Answer Sheet To Chemistry 1 Quiz Review Stoichiometry eBook Subscription Services
 - Answer Sheet To Chemistry 1 Quiz Review Stoichiometry Budget-Friendly Options
6. Navigating Answer Sheet To Chemistry 1 Quiz Review Stoichiometry eBook Formats
- ePub, PDF, MOBI, and More
 - Answer Sheet To Chemistry 1 Quiz Review Stoichiometry Compatibility with Devices
 - Answer Sheet To Chemistry 1 Quiz Review Stoichiometry Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
 - Highlighting and Note-Taking Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
 - Interactive Elements Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
8. Staying Engaged with Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
9. Balancing eBooks and Physical Books Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
- Setting Reading Goals Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
- Fact-Checking eBook Content of Answer Sheet To Chemistry 1 Quiz Review Stoichiometry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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