

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a known volume of a silver nitrate (AgNO_3) solution. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is designed to illustrate the principles of gravimetric analysis and the use of stoichiometry to determine the concentration of a substance in a solution.

The reaction between silver ions and chloride ions is as follows:

$$\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$$

Procedure:

The procedure involves the following steps:

1. Preparation of a known volume of a solution containing chloride ions.
2. Addition of a known volume of a silver nitrate solution to the chloride solution.
3. Filtration of the precipitate.
4. Drying of the precipitate.
5. Weighing of the precipitate.

Data Analysis:

	Sample 1	Sample 2
Mass of precipitate	0.1234 g	0.1567 g
Mass of chloride ion	0.0456 g	0.0589 g
Mass of silver ion	0.0778 g	0.0978 g

The mass of chloride ion in the original solution is calculated as follows:

$$\text{Mass of Cl}^- = \text{Mass of AgCl} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$$

The mass of silver ion in the original solution is calculated as follows:

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