Anions, Cations, and Ionic Reactions

Hands-On Labs, Inc. Version 42-0304-00-01

Lab Report Assistant

This document is not meant to be a substitute for a formal laboratory report. The Lab Report Assistant is simply a summary of the experiment's questions, diagrams if needed, and data tables that should be addressed in a formal lab report. The intent is to facilitate students' writing of lab reports by providing this information in an editable file which can be sent to an instructor.

Exercise 1: Qualitative Anion Tests

Data Table 1. Anion Confirmation Tests.

Chemical	Anion	Observations: Addition of AgNO ₃	Observations: Addition of HCI
NaBr	Bromide	Turned milky white	Turned yellowish More of a gel
Na ₂ CO ₃	Carbonate	Thick white glob A little grainy	White sediment in a clear liquid Bubbles up
NaCl	Chloride	Milky white	White sediment visible
Nal	lodide	Yellow glob immediately when dropped in	Yellow, grainy/sediment, thick looking
Na ₂ SO ₄	Sulfate	Thickened	Very cloudy/milky
Na₂S	Sulfide	Brown Globs immediately when dropped in Still watery	Brown globs thinned out into brown pieces of sediment Still watery (brownish) looks like dirty water
Unknown #103	Carbonate	Thick white glob Turned gel	Bubbly and cloudy Very thick gel
Unknown #104	Sulfide	Brown globs immediately when dropped in Still watery	Brown globs thinned out into brown pieces of sediment Still watery (brownish) looks like dirty water

Analysis Of Anions Cations Lab Answers

J Dewey

Analysis Of Anions Cations Lab Answers:

CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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Electrochemistry Mohammed Khalid,2013-02-20 Electrochemistry has been undergoing significant transformations in the last few decades It is now the province of academics interested only in measuring thermodynamic properties of solutions and of industrialists using electrolysis or manufacturing batteries with a huge gap between them It has become clear that these apparently distinct subjects alongside others have a common ground and that they have grown towards each other particularly as a result of research into the rates of electrochemical processes Such evolution is due to a number of factors and offers the possibility of carrying out reproducible dynamic experiments under an ever increasing variety of conditions with reliable and sensitive instrumentation This has enabled many studies of a fundamental and applied nature to be carried out **Geological Survey Water-supply Paper** ,1985 The Office of Environmental Management Technical Reports ,1997

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