

Solution

1. Answer (A, B, D)

n. factor of KMnO_4 in acidic medium = 5

In faintly basic/neutral medium = 3

2. Answer (A, D)

n.f. for $\text{C}_2\text{O}_4^{2-}$ or HC_2O_4^- with $\text{K}_2\text{Cr}_2\text{O}_7$ = 2

n.f. for HC_2O_4^- with HCl = 1 and $\text{C}_2\text{O}_4^{2-}$ with HCl = 2

3. Answer (C)



4. Answer (C)

Let the two metal carbonates be MCO_3 and $\text{M}'\text{CO}_3$. Let M be 13.6% by weight.

Weight of sample = 2.5 g

Weight of metal M = $\frac{13.6}{100} \times 2.5 = 0.34 \text{ g}$

Weight of CO_2 evolved = 1.33g

Moles of CO_2 evolved = $\frac{1.33}{44} = 0.0302$

$\therefore$ Weight of CO_3^{2-} in the sample = $0.0302 \times 60 = 1.812 \text{ g}$

$\therefore$ % of M' = $\frac{2.5 - 0.34 - 1.812}{2.5} \times 100 = 13.92\%$

5. Answer (A, B, C)



Eq. of H_2S = equivalent of I_2

$$\Rightarrow n_{\text{H}_2\text{S}} \times 2 = 2.40 \times 0.02$$

$$\Rightarrow n_{\text{H}_2\text{S}} = 2.40 \times 10^{-2} \text{ m moles}$$

$\therefore$ % of S in steel sample

$$= \frac{2.4 \times 10^{-3} \times 32}{0.3 \times 100} \times 100$$

$$= 0.256\%$$

6. Answer (A, B, D)

For equimolar mixture of H_2SO_4 and SO_3

$$x + 80 + x \times 98 = 10$$

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 - The Rise of Digital Reading 2008 Chemistry Sol Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 2008 Chemistry Sol Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2008 Chemistry Sol Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2008 Chemistry Sol Answers
 - Personalized Recommendations
 - 2008 Chemistry Sol Answers User Reviews and Ratings
 - 2008 Chemistry Sol Answers and Bestseller Lists
5. Accessing 2008 Chemistry Sol Answers Free and Paid eBooks
 - 2008 Chemistry Sol Answers Public Domain eBooks
 - 2008 Chemistry Sol Answers eBook Subscription Services
 - 2008 Chemistry Sol Answers Budget-Friendly Options

6. Navigating 2008 Chemistry Sol Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 2008 Chemistry Sol Answers Compatibility with Devices
 - 2008 Chemistry Sol Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2008 Chemistry Sol Answers
 - Highlighting and Note-Taking 2008 Chemistry Sol Answers
 - Interactive Elements 2008 Chemistry Sol Answers
8. Staying Engaged with 2008 Chemistry Sol Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2008 Chemistry Sol Answers
9. Balancing eBooks and Physical Books 2008 Chemistry Sol Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2008 Chemistry Sol Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2008 Chemistry Sol Answers
 - Setting Reading Goals 2008 Chemistry Sol Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2008 Chemistry Sol Answers
 - Fact-Checking eBook Content of 2008 Chemistry Sol Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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