

Practice Test Ch 3

Stoichiometry

Name: _____ Per: _____

- Remember, this is practice - Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math" – common factors and rough estimation – do not do "long division" to try to get exact values. Remember it is a MC test, use the answers
- Mark which questions you would like to "go over" when we get to school in September.

1. Balance the following equation:



The balanced equation shows that 1.00 mole of NH_3 requires _____ mole(s) of O_2 :

- a. 0.57
 - b. 1.25
 - c. 1.33
 - d. 1.75
 - e. 3.5
2. Write a balanced equation for the combustion of acetaldehyde, CH_3CHO .
- When properly balanced, the equation indicates that _____ mole(s) of O_2 are required for each mole of CH_3CHO .
- a. 1
 - b. 2
 - c. 2.5
 - d. 3
 - e. 5
3. What is the total mass of products formed when 16 grams of CH_4 is burned with excess oxygen?
- a. 32 g
 - b. 36 g
 - c. 44 g
 - d. 62 g
 - e. 80 g
4. Write a balanced equation for the combustion of propane, C_3H_8 .

When balanced, the equation indicates that _____ moles of O_2 are required for each mole of C_3H_8 .

- a. 1.5
- b. 3
- c. 3.5
- d. 5
- e. 8

5. Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the
- sum
- of the coefficients in the balanced equation:



- a. 3
 - b. 5
 - c. 6
 - d. 7
 - e. 8
6. Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below.
- $$2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$$
- a. 1.5 g
 - b. 2.0 g
 - c. 3.0 g
 - d. 6.0 g
 - e. 12 g
7. How many grams of nitric acid, HNO_3 , can be prepared from the reaction of 128 g of NO_2 with 54.0 g H_2O according to the equation below?
- $$3\text{NO}_2 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3 + \text{NO}$$
- a. 92
 - b. 108
 - c. 126
 - d. 189
 - e. 279
8. Which of the following statements is true?
- I. The molar mass of CaCO_3 is 100.1 g mol^{-1} .
 - II. 50 g of CaCO_3 contains 9×10^{23} oxygen atoms.
 - III. A 200 g sample of CaCO_3 contains 2 moles of CaCO_3 .

- a. I only
- b. II only
- c. III only
- d. I and III only
- e. I, II, and III

Answer Key To Stoichiometry Test Ch 1

Arshad Iqbal



Answer Key To Stoichiometry Test Ch 1:

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Instructor's Manual and Test Bank to Accompany Basic Concepts of Chemistry Malone, 2000-04-18

Chemistry for the Utterly Confused John T. Moore, Richard H. Langley, 2007-06-01 Banish bafflement in this tough subject From formulas and lab techniques to the periodic table Chemistry for the Utterly Confused focuses on the areas of maximum confusion and breaks down the most difficult chemistry topics into easy to understand concepts This invaluable

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