

Ternary compounds have \_\_\_\_\_ (three) \_\_\_\_\_ (three or) elements. They are usually always composed of a metal/oid \_\_\_\_\_ (metalloid) nonmetal and a polyatomic \_\_\_\_\_ (nonmetal) anion. The only positively charged polyatomic ion is \_\_\_\_\_ (ammonium). To write the formula of a ternary compound is no different than to write the formula of a binary compound with one exception. If a subscript is necessary for the polyatomic ion in order to \_\_\_\_\_ (balance) charge, we must place the polyatomic ion in \_\_\_\_\_ (parentheses).

As each ion, write the formula of the ionic compound consisting of the cation on the left of the dash and the negative ion on the right of the dash.

|                  | $\text{OH}^-$ | $\text{SO}_4^{2-}$ | $\text{NO}_3^-$ | $\text{CO}_3^{2-}$ | $\text{PO}_4^{3-}$ |
|------------------|---------------|--------------------|-----------------|--------------------|--------------------|
| $\text{K}^+$     | _____         | _____              | _____           | _____              | _____              |
| $\text{Na}^{2+}$ | _____         | _____              | _____           | _____              | _____              |
| $\text{NH}_4^+$  | _____         | _____              | _____           | _____              | _____              |
| $\text{Al}^{3+}$ | _____         | _____              | _____           | _____              | _____              |
| $\text{Ca}^{2+}$ | _____         | _____              | _____           | _____              | _____              |
| $\text{Fe}^{2+}$ | _____         | _____              | _____           | _____              | _____              |

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**Karin Nielsen-Saines**



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