

Chemistry Final Exam Review Questions

Useful Values and Equations

Some of the following values and equations may be useful for answering some of the questions and problems.

$$Q_F = 1.84 Q_C + 12$$

speed of light = 3.00×10^8 m/s

2000

$R = 0.082057 \text{ L-atm / deg-mole}$

Planck's constant = $6.63 \times 10^{-34} \text{ J.s}$

15. *Journal of the American Medical Association*, 2000; 284: 1039-1044.

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- One of the following does **not** describe gases. Which one is it?
 - Expands infinitely.
 - Rigid shape.
 - Easily compressed.
 - Relatively low densities.
- The following properties describe the element, mercury (Hg). Which one is a **chemical** property?
 - Its density is 5.43 g/cm^3 .
 - It is a silvery-white metal.
 - It is a liquid at 298 K.
 - Mercury(II) sulfide forms when it is exposed to sulfide ions.
- All of the following are properties of SO_2 . Which one is a **physical** property?
 - It reacts with oxygen to form SO_3^{2-} .
 - It condenses to a colorless liquid at 14°F and atmospheric pressure.
 - It is a strong reducing agent.
 - It reacts with water and oxygen to form an acid.
- Which of the following numbers has 4 significant figures?
 - 0.04309
 - 0.0430
 - 0.0431
 - 0.43980
- How many significant figures should there be in the answer to the following math equation?
$$2.834 + 5.71520 + 2.12 + 178.1 + 250.2619$$
 - 3
 - 4
 - 5
 - 6
- How many significant figures should there be in the answer to the following math equation?
$$23.413 \text{ g} + (2.15 \text{ cm} \times 1.1 \text{ cm} \times 3.73 \text{ cm})$$
 - 2
 - 3
 - 4
 - 5
- At 25°C , one cm^3 of iron has a mass of 7.62 grams. What volume of iron is occupied by 125 grams of iron?
 - 0.82 cm^3
 - 8.2 cm^3
 - 0.164 cm^3
 - 16.4 cm^3
- A cube of lead has a volume of 6.61 cm^3 . Its mass is 75.0 grams. What is its density?
 - 3.85 g/cm^3
 - 7.41 g/cm^3
 - 11.3 g/cm^3
 - 22.6 g/cm^3

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web aug 19 2022 atoms with the same number of protons but different numbers of neutrons are called isotopes they share almost the same chemical properties but differ in mass and therefore in physical properties there are stable isotopes which do not emit radiation and there are unstable isotopes which do emit radiation

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usually one or two isotopes of an element are the most stable and common different isotopes of an element generally have the same physical and chemical properties because they have the same numbers of protons and electrons

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web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and position in the periodic table and nearly identical chemical behavior but with different atomic masses and physical properties every chemical element has one or more isotopes

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