

SCH4U Unit 4 - Quiz – Equilibrium

Player: _____ Period: _____

120 130 140 150

160 170

180

190 200

210

220 230 240 250

$$K_{eq} = \frac{[products]}{[reactants]} \quad K_{eq} = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

$$K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4}^2 P_{O_2}} \quad K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{(P_{CH_4})^2 (P_{O_2})}$$

$$pH = -\log[H^+] \quad pOH = -\log[OH^-] \quad pK_a = -\log(K_a) \quad K_a \times K_b = K_w$$

Question 1:

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 100°C, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , and 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.020 atm. The reaction is:



If the volume of an 8.00 L container at 100°C where a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
 Answer: $[H_2] = 0.0025 \text{ mol/L}$
 Answer: $[H_2] = 0.020 \text{ mol/L}$
 $[H_2] = 0.010 \text{ mol/L}$
- Determine the value of K_{eq} at equilibrium.
- Answer:
 What other equilibria have associated formula constants at 100°C involving H_2 and H_2S ?
- Calculate:

	$K_{eq} = \frac{[H_2S]}{[H_2]}$	$K_{eq} = \frac{[H_2S]}{[H_2]}$	$K_{eq} = \frac{[H_2S]}{[H_2]}$
Initial concentration (mol/L)			
Change in concentration (mol/L)			
Final concentration (mol/L)	0.020	0.020	0.020

What is the final concentration?

Answers For Sch4u Unit 4

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This book delves into Answers For Sch4u Unit 4. Answers For Sch4u Unit 4 is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Answers For Sch4u Unit 4, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Answers For Sch4u Unit 4
- Chapter 2: Essential Elements of Answers For Sch4u Unit 4
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- Chapter 4: Answers For Sch4u Unit 4 in Specific Contexts
- Chapter 5: Conclusion

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Table of Contents Answers For Sch4u Unit 4

1. Understanding the eBook Answers For Sch4u Unit 4

- The Rise of Digital Reading Answers For Sch4u Unit 4
- Advantages of eBooks Over Traditional Books
- 2. Identifying Answers For Sch4u Unit 4
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Sch4u Unit 4
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answers For Sch4u Unit 4
 - Personalized Recommendations
 - Answers For Sch4u Unit 4 User Reviews and Ratings
 - Answers For Sch4u Unit 4 and Bestseller Lists
- 5. Accessing Answers For Sch4u Unit 4 Free and Paid eBooks
 - Answers For Sch4u Unit 4 Public Domain eBooks
 - Answers For Sch4u Unit 4 eBook Subscription Services
 - Answers For Sch4u Unit 4 Budget-Friendly Options
- 6. Navigating Answers For Sch4u Unit 4 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Sch4u Unit 4 Compatibility with Devices
 - Answers For Sch4u Unit 4 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Sch4u Unit 4
 - Highlighting and Note-Taking Answers For Sch4u Unit 4
 - Interactive Elements Answers For Sch4u Unit 4
- 8. Staying Engaged with Answers For Sch4u Unit 4
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Sch4u Unit 4

9. Balancing eBooks and Physical Books Answers For Sch4u Unit 4
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Sch4u Unit 4
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Sch4u Unit 4
 - Setting Reading Goals Answers For Sch4u Unit 4
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Sch4u Unit 4
 - Fact-Checking eBook Content of Answers For Sch4u Unit 4
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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