

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Sub-levels have different amounts of energy. For example, orbitals closer to the nucleus have more energy than orbitals further from the nucleus. The following orbitals have different levels of energy:



For each pair, write sublevel(s) for which orbitals are included. Then draw orbital(s) representing an orbital.



How does it work? Orbitals appear in the order of their energy. The lowest energy sublevel is filled first. For example, the 4s sublevel has a higher energy than a 3d sublevel, so the 4s sublevel has a higher energy than the 3d sublevel.

When electrons occupy orbitals, they try to have the lowest possible energy possible. (This is called the **lowest energy principle**.) An electron will occupy a sublevel before the next level is filled. For example, the 4s sublevel is filled before the 3d sublevel is filled. Additionally, the 4s sublevel is filled before the 3d sublevel is filled. (This is called the **Pauli Exclusion Principle**.) When two electrons occupy the same orbital, they must spin in opposite directions—one clockwise and the other counter-clockwise.

Critical Thinking Questions

- a. How many electrons would it take to fill a sublevel before beginning to fill the next? b. How many electrons would it take to fill a sublevel before beginning to fill the next? c. How many electrons would it take to fill a sublevel before beginning to fill the next? d. How many electrons would it take to fill a sublevel before beginning to fill the next? e. How many electrons would it take to fill a sublevel before beginning to fill the next? f. How many electrons would it take to fill a sublevel before beginning to fill the next? g. How many electrons would it take to fill a sublevel before beginning to fill the next? h. How many electrons would it take to fill a sublevel before beginning to fill the next? i. How many electrons would it take to fill a sublevel before beginning to fill the next? j. How many electrons would it take to fill a sublevel before beginning to fill the next? k. How many electrons would it take to fill a sublevel before beginning to fill the next? l. How many electrons would it take to fill a sublevel before beginning to fill the next? m. How many electrons would it take to fill a sublevel before beginning to fill the next? n. How many electrons would it take to fill a sublevel before beginning to fill the next? o. How many electrons would it take to fill a sublevel before beginning to fill the next? p. How many electrons would it take to fill a sublevel before beginning to fill the next? q. How many electrons would it take to fill a sublevel before beginning to fill the next? r. How many electrons would it take to fill a sublevel before beginning to fill the next? s. How many electrons would it take to fill a sublevel before beginning to fill the next? t. How many electrons would it take to fill a sublevel before beginning to fill the next? u. How many electrons would it take to fill a sublevel before beginning to fill the next? v. How many electrons would it take to fill a sublevel before beginning to fill the next? w. How many electrons would it take to fill a sublevel before beginning to fill the next? x. How many electrons would it take to fill a sublevel before beginning to fill the next? y. How many electrons would it take to fill a sublevel before beginning to fill the next? z. How many electrons would it take to fill a sublevel before beginning to fill the next?

Information: Hund's Rule

Electrons prefer to "spin" or "upspin". Electrons prefer to spin in the same direction. If two electrons occupy the same orbital, they must spin in opposite directions. (This is called the **Pauli Exclusion Principle**.)



If two electrons occupy the same orbital, they must spin in opposite directions. (This is called the **Pauli Exclusion Principle**.) If two electrons occupy the same orbital, they must spin in opposite directions. (This is called the **Pauli Exclusion Principle**.)



Hund's rule says to fill orbitals in the same direction. Hund's rule says to fill orbitals in the same direction. (This is called the **Pauli Exclusion Principle**.) Hund's rule says to fill orbitals in the same direction. (This is called the **Pauli Exclusion Principle**.)

The following are the electron orbital diagrams for the elements, nitrogen and oxygen. Write the element's 2p sublevel and explain to show Hund's rule. The 2p sublevel is filled before the 3s sublevel is filled.



Critical Thinking Questions

- a. How many "upspin" electrons are in a nitrogen atom? b. How many "upspin" electrons are in a nitrogen atom? c. How many "upspin" electrons are in a nitrogen atom? d. How many "upspin" electrons are in a nitrogen atom? e. How many "upspin" electrons are in a nitrogen atom? f. How many "upspin" electrons are in a nitrogen atom? g. How many "upspin" electrons are in a nitrogen atom? h. How many "upspin" electrons are in a nitrogen atom? i. How many "upspin" electrons are in a nitrogen atom? j. How many "upspin" electrons are in a nitrogen atom? k. How many "upspin" electrons are in a nitrogen atom? l. How many "upspin" electrons are in a nitrogen atom? m. How many "upspin" electrons are in a nitrogen atom? n. How many "upspin" electrons are in a nitrogen atom? o. How many "upspin" electrons are in a nitrogen atom? p. How many "upspin" electrons are in a nitrogen atom? q. How many "upspin" electrons are in a nitrogen atom? r. How many "upspin" electrons are in a nitrogen atom? s. How many "upspin" electrons are in a nitrogen atom? t. How many "upspin" electrons are in a nitrogen atom? u. How many "upspin" electrons are in a nitrogen atom? v. How many "upspin" electrons are in a nitrogen atom? w. How many "upspin" electrons are in a nitrogen atom? x. How many "upspin" electrons are in a nitrogen atom? y. How many "upspin" electrons are in a nitrogen atom? z. How many "upspin" electrons are in a nitrogen atom?

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