

1 Complete the word equations for the reactions of the Group 7 elements with metals.

- a iron + chlorine → iron chloride  
 b iron + bromine → iron bromide  
 c sodium + iodine → sodium iodide  
 d potassium + bromine → potassium bromide  
 e lithium + chlorine → lithium chloride  
 f sodium + chlorine → sodium chloride

2 Make up four sentences using the phrases in the table. Each sentence must include a phrase from each column.

|           |               |                                                 |
|-----------|---------------|-------------------------------------------------|
| There are | weak forces   | between the two atoms of a chlorine molecule.   |
|           | strong forces | between a chlorine molecule and its neighbours. |
|           |               | between the two atoms of an iodine molecule.    |
|           |               | between an iodine molecule and its neighbours.  |

*There are weak forces between the two atoms of a chlorine molecule.  
 There are strong forces between a chlorine molecule and its neighbours.  
 There are weak forces between the two atoms of an iodine molecule.  
 There are strong forces between an iodine molecule and its neighbours.*

3 For the sentences A–H below:

- Write **1** next to the sentences that are true for Group 1.
  - Write **7** next to the sentences that are true for Group 7.
  - Write **B** next to the sentences that are true for both Groups 1 and 7.
- A Going down this group, boiling point increases.  
 B Going up this group, the reactions of the elements get more vigorous.  
 C Going down this group, the number of protons in atoms of the elements increases.  
 D Atoms of the elements in this group form molecules made up of two atoms.  
 E Going up this group, the colours of the elements get lighter.  
 F The elements at the top of this group are gases.  
 G The elements in this group conduct electricity.  
 H All the atoms of the elements in this group have the same number of electrons in the outermost shell.

*7*  
*1*  
*B*  
*1*  
*B*  
*7*  
*B*

12 Write **M** next to the pair of elements that you predict will react most vigorously.  
 Write **L** next to the pair of elements that you predict will react least vigorously.

- a sodium and iodine  
 b lithium and iodine  
 c potassium and fluorine  
 d potassium and bromine  
 e sodium and bromine

*L*  
*L*  
*M*  
*M*  
*L*

# Answers For A Ls Chemistry

**Charles E. Mortimer**



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