

General Instruction:

- This question paper consists of 39 questions in 5 sections.
- All questions are compulsory. However, an internal choice is provided in some questions. A student is expected to attempt only one of these questions.
- Section A consists of 20 objective type questions carrying 1 mark each.
- Section B consists of 6 Very Short questions carrying 02 marks each. Answers to these questions should be in the range of 30 to 50 words.
- Section C consists of 7 Short Answer type questions carrying 03 marks each. Answers to these questions should be in the range of 50 to 80 words.
- Section D consists of 3 Long Answer type questions carrying 05 marks each. Answer to these questions should be in the range of 80 to 120 words.
- Section E consists of 3 source-based/case-based units of assessment of 04 marks each with sub-parts

SECTION – A

- MCQs: [1 Mark Each]

[20]

- Which one of the following statements is not true?
 - The molecules in a solid vibrate about a fixed position.
 - The molecules in a liquid are arranged in a regular pattern.**
 - The molecules in a gas exert negligibly small forces on each other, except during collisions.
 - The molecules of a gas occupy all the space available.
- When a gas jar containing colourless air is kept upside down over a gas jar full of brown coloured bromine vapour, then after some time, the brown colour of bromine vapour spreads into the upper jar making both the gas jar appear brown in colour. Which of the following conclusions obtained from these observations is incorrect?
 - Bromine vapour is made of tiny particles which are moving.
 - Air is made up of tiny particles which are moving.
 - The particles of bromine are moving but those of air are not moving.**
 - Even though bromine vapour is heavier than air, it can move up against gravity.
- A mixture of sulphur and carbon disulphide is:
 - Heterogeneous and shows Tyndall effect
 - Homogeneous and shows Tyndall effect
 - Heterogeneous and does not show Tyndall effect
 - Homogeneous and does not show Tyndall effect**
- Tincture of Iodine has antiseptic properties. This solution is made by dissolving:
 - Iodine in potassium iodide
 - Iodine in acetone
 - Iodine in water
 - Iodine in alcohol**
- If 1.4 g of calcium oxide is formed by the complete decomposition of calcium carbonate, then the amount of calcium carbonate taken and the amount of carbon dioxide formed will be respectively:

CaCO_3
 Calcium Carbonate

$\xrightarrow{\text{Heat}}$

CaO
 Calcium oxide

$+$

CO_2
 Carbon dioxide

 - 2.2 g and 1.1 g
 - 1.1 g and 2.5 g
 - 2.5 g and 1.1 g**
 - 5.0 g and 1.1 g

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SA Adler



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