

Analysis of Alum $\text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$

After a compound has been synthesized, tests should be carried out to verify the compound formed is indeed the compound desired. There are a number of various tests that can be performed to verify that the compound is the one desired. In a previous experiment, alum crystals were prepared. In this experiment we will do several tests to determine if the crystals are really alum.

The first and simplest test is to find the melting point of the compound and compare it to the published value for alum. A small quantity of alum is powdered and placed on a melting point apparatus. The crystals are heated on the melting point apparatus on a hot plate, and the temperature at which they melt is recorded and compared to the reported values.

The second test that we can do is to determine the amount of water of hydration present in the alum crystals. Some of the alum is placed in a crucible and weighed. The crucible is heated until all of the water of hydration is driven off. The crucible is then cooled and its mass measured. From the mass of the dry crystals and the mass of the water lost, the ratio of moles water to moles $\text{KAl}(\text{SO}_4)_2$ can be calculated and then compared to the correct formula values.

The third test is a chemical test to determine the percent of sulfate in the compound prepared. A weighed quantity of alum is dissolved in distilled water. An excess of barium ions is added to the solution to precipitate all of the sulfate in the compound as barium sulfate. The precipitated barium sulfate is filtered off, dried, and its mass determined. From the mass of the barium sulfate and the initial mass of the alum, the percent sulfate can be calculated and compared with the theoretical percent found from the formula.

Chemicals

Parts 1, 2, and 3

Alum crystals prepared by you in the Lab
Known Alum crystals

Part 3

Barium nitrate, $\text{Ba}(\text{NO}_3)_2$, 0.2M

Equipment

Part 1

Capillary Tube	Thermometer
Melting point apparatus	Mortar and pestle
Hot plate	

Part 2

Crucible and cover	Ring stand, ring and Bunsen burner
Clay triangle	

Analysis Of Alum Lab Answers Sulfate

Victor M. Corman



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Table of Contents Analysis Of Alum Lab Answers Sulfate

1. Understanding the eBook Analysis Of Alum Lab Answers Sulfate
 - The Rise of Digital Reading Analysis Of Alum Lab Answers Sulfate
 - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of Alum Lab Answers Sulfate
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analysis Of Alum Lab Answers Sulfate
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of Alum Lab Answers Sulfate
 - Personalized Recommendations
 - Analysis Of Alum Lab Answers Sulfate User Reviews and Ratings
 - Analysis Of Alum Lab Answers Sulfate and Bestseller Lists
5. Accessing Analysis Of Alum Lab Answers Sulfate Free and Paid eBooks

- Analysis Of Alum Lab Answers Sulfate Public Domain eBooks
- Analysis Of Alum Lab Answers Sulfate eBook Subscription Services
- Analysis Of Alum Lab Answers Sulfate Budget-Friendly Options
- 6. Navigating Analysis Of Alum Lab Answers Sulfate eBook Formats
 - ePub, PDF, MOBI, and More
 - Analysis Of Alum Lab Answers Sulfate Compatibility with Devices
 - Analysis Of Alum Lab Answers Sulfate Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analysis Of Alum Lab Answers Sulfate
 - Highlighting and Note-Taking Analysis Of Alum Lab Answers Sulfate
 - Interactive Elements Analysis Of Alum Lab Answers Sulfate
- 8. Staying Engaged with Analysis Of Alum Lab Answers Sulfate
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analysis Of Alum Lab Answers Sulfate
- 9. Balancing eBooks and Physical Books Analysis Of Alum Lab Answers Sulfate
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analysis Of Alum Lab Answers Sulfate
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Analysis Of Alum Lab Answers Sulfate
 - Setting Reading Goals Analysis Of Alum Lab Answers Sulfate
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analysis Of Alum Lab Answers Sulfate
 - Fact-Checking eBook Content of Analysis Of Alum Lab Answers Sulfate
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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