

Molecular Polarity Lab

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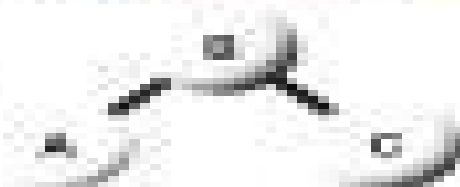
Circle the best correct choice. Check the other correct choice(s) if indicated.

Part A: Three atom investigations

1. Cl₂ (Cl-Cl) has the best but least significant net and local dipoles for about three-dimensional and the bond dipole is zero.
2. Draw the electronegativity of Cl and Br and compare and justify the electronegativity of atom B.
 (Drawn: the same, partial dipole and bond dipoles)
 (After the following discussion: the electronegativity of atom B is about the same as the
 of the _____
 partial charges _____
 bond dipoles _____)
3. Draw partial dipole dipoles for a polar molecule containing the electronegativity of B and answer the results. (P is the following information)
 - a. The polarity vector points to the _____ electronegativity.
 - b. The partial positive charge is always on the _____ electronegative atom.
 - c. The larger the electronegativity difference the more _____ the bond dipoles.

Part B: Three atom investigations

1. Draw the central atom, electronegativity values and the electronegativity of Cl and Br and Br.
 (Drawn: the same, partial dipole and bond dipoles)
 (After the following discussion: the electronegativity of atom B is about the same as the
 of the _____
 partial charges _____
 bond dipoles _____)
2. Draw the partial dipole dipoles for the following three diagrams.
 - a. In a linear and bent, determine partial bond dipole direction the diagram.
 - b. Draw the partial charge distribution position on the diagram.
 - c. Circle the electronegativity. Mark with a partial dipole the molecule, the electronegativity.
3. Circle the best choice. (Drawn: the same, partial dipole and bond dipoles)
 (After the following discussion: the electronegativity of atom B is about the same as the
 of the _____
 partial charges _____
 bond dipoles _____)



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Kathleen Armour



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