

SCH4U Unit 4 - Quiz – Equilibrium

Name: _____ Period: _____

1201-1202

1203

1204

1205

1206

Write the formulas for:

$$K_{eq} = \frac{[\text{products}]}{[\text{reactants}]} \quad K_{eq} = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

$$K_{eq} = \frac{P_{\text{H}_2\text{O}}^2 P_{\text{CO}_2}}{P_{\text{CH}_4} P_{\text{CO}}^2} \quad K_{eq} = \frac{P_{\text{H}_2\text{O}}^2 P_{\text{CO}_2}}{P_{\text{CH}_4} P_{\text{CO}}^2}$$

gaseous H_2O (g) + gaseous CO_2 (g) + gaseous CH_4 (g) + gaseous CO (g)

Question 1

1. Hydrogen sulfide, H_2S , is a gaseous compound. At 1000°C, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, H_2S , and an undetermined amount of sulfur, S_2 . The reaction is:



If the volume of the 1.000-L container is 1.000 L, what concentration of S_2 is present at equilibrium?

- What is the concentration of H_2 at equilibrium?
- What is the concentration of H_2S at equilibrium?
- What is the concentration of S_2 at equilibrium?
- What is the concentration of S_2 at equilibrium?
- What is the concentration of S_2 at equilibrium?
- What is the concentration of S_2 at equilibrium?
- What is the concentration of S_2 at equilibrium?
- What is the concentration of S_2 at equilibrium?

	H_2	H_2S	S_2
Initial (moles)			
Change (moles)			
Equilibrium (moles)			

What is the concentration of S_2 at equilibrium?

Answers For Sch4u Unit 4

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