

## Chemistry Practice Exam v2

**Directions:** It is important that you provide answers in your own words. Please focus only on information from the text/eBook to create your own solutions. Please do not use direct information from an outside source (especially copying and pasting from an "answer" website). Use of direct information from an outside source is against school policy. *All answers will be checked for plagiarism. Instances of plagiarism can result in probation or possible dismissal from the school.*

Please be sure to follow all guidelines (number of sentences/showing all calculations) and to provide the correct metric units of measure. All questions are 3.33 points each (.33 for unit)

1. A student collected the following data for a fixed volume of gas:

| Temperature ( $^{\circ}\text{C}$ ) | Pressure (mm of Hg) |
|------------------------------------|---------------------|
| 10                                 | 726                 |
| 20                                 | 750                 |
| 40                                 | 800                 |
| 70                                 | 880                 |
| 100                                | 960                 |
| 150                                | 1085                |

Fill in the missing data point. Show all calculations leading to an answer.

The General Law of Gases states that

$$P_1 \cdot V_1 / T_1 = P_2 \cdot V_2 / T_2.$$

As Volume is constant, we can write this as

$$P_1 / T_1 = P_2 / T_2.$$

$P_1$  and  $T_1$  are the pressure and temperature at time 1.

$P_2$  and  $T_2$  are the pressure and temperature at time 2.

You can take any pair of values as  $P_1$  and  $T_1$  and find  $P_2 = P_1 T_2 / T_1$ , once you sum 273.15 to each temperature to get kelvins.

$$P_1 = 726 \text{ mm Hg}$$

$$T_1 = 10 + 273.15 = 283.15 \text{ K};$$

$$T_2 = 150 + 273.15 = 423.15 \text{ K}.$$

$$\text{Then, } P_2 = 726 \cdot 423.15 / 283.15 = 1085.96 \text{ mm Hg}.$$

2. You are given a clear solution of  $\text{KNO}_3$ . Using 3 – 4 sentences (in your own words) explain how you would determine if the solution is unsaturated, saturated or supersaturated.

# Answers Modeling Chemistry U7 Empfrm Practice V2 0

**Clifford Lane**



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