

## Solutions

| <b>Word</b>          | <b>Definition</b>                                                                                                                                                                                                                                            |
|----------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| Aqueous              | A solution in which the solvent is water.                                                                                                                                                                                                                    |
| Colligative property | A property of a solution that is dependent on concentration. Examples include boiling point, freezing point and vapor pressure.                                                                                                                              |
| Molality             | The concentration of a solution measured in moles of solute per kilogram of solvent.                                                                                                                                                                         |
| Molarity             | The concentration of a solution measured in moles of solute per liter of solution.                                                                                                                                                                           |
| Parts per million    | The concentration of a solution measured in mass of solute per mass of solution multiplied by one million.                                                                                                                                                   |
| Percent by mass      | The concentration of a solution measured in mass of solute per mass of solution multiplied by one hundred.                                                                                                                                                   |
| Percent by volume    | The concentration of a solution measured in volume of solute per volume of solution multiplied by one hundred.                                                                                                                                               |
| Saturated            | A solution that has the maximum concentration of solute possible in a given quantity of solvent at a given temperature, a solution at equilibrium.                                                                                                           |
| Solubility           | The maximum quantity of solute that can be dissolved in a given quantity of solvent at a given temperature to make a saturated solution.                                                                                                                     |
| Solute               | A substance that is torn apart and kept separate by solvent particles.                                                                                                                                                                                       |
| Solution             | A homogeneous mixture formed from a solute dissolved into a solvent.                                                                                                                                                                                         |
| Solvent              | A substance that attaches to solute particles, tears them apart from each other and keeps them apart.                                                                                                                                                        |
| Supersaturated       | A solution in which there is an excess of solute beyond the solubility point for a given temperature. Either the excess will precipitate or it will remain precariously dissolved until the solution is agitated, whereupon the excess will precipitate out. |
| Unsaturated          | A solution in which there are still solvent particles free to attach to and dissolve more solute particles.                                                                                                                                                  |

# 2011 Mark H Rosengarten Unit 10 Answers

**Ying-Ying Zheng**



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