

Antacid Analysis and Titration – Lab Report Assistant

Exercise 1: Back Titration of Antacid Neutralization

Data Table 1. Antacid Neutralization Data

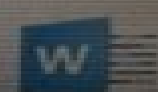
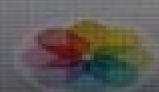
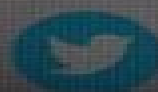
Mass of Crushed Antacid (g)	
Concentration of HCl (M)	1M
Volume HCl (mL)	5 mL
Concentration of NaOH (M)	1M
Initial NaOH Volume (mL)	
Final NaOH Volume (mL)	
Total Volume of NaOH Used (mL)	

Data Table 2. Experimental Results

HCl available for neutralization (g):	0.3646 (g)
Moles of NaOH required to reach stoichiometric point:	0.0039 (moles)
HCl neutralized by antacid:	0.222 (g)
HCl neutralized per gram of antacid:	0.445 (g)

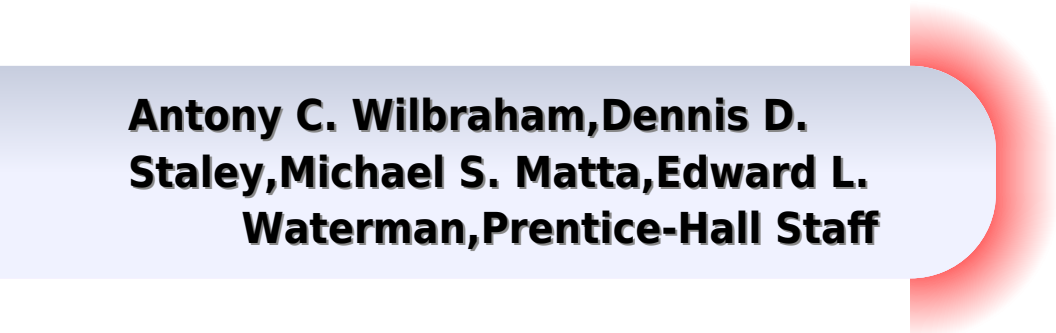
Questions:

- If an antacid tablet weighed 1.6 grams, how many moles of gastric acid (HCl) would it neutralize? Use the results obtained in **Data Tables 1 and 2** to explain and quantify your answer.
- If you performed this experiment with a titrant of 0.5M NaOH, would you expect your results in **Data Table 2** to change or stay the same? Explain your answer.
- The reaction that occurred when the antacid mixed with the HCl resulted in an additional product (besides a salt and water). Did you see evidence of this product? Describe the experimental evidence you witnessed that supports the formation of the additional product.



Analysis Of Antacid Manual Answers

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Waterman, Prentice-Hall Staff**



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