

Name: _____

Sci # _____

COVALENT BOND WEBQUEST

DIRECTIONS: Answer the following by going to the notes website. You can go to Mrs. Slowiak's website and use the links or type in the addresses for each section if needed.

Activity #1- Introduction to Covalent Bonding

Open [Chemical Bonding](http://www.visionlearning.com/library/module_viewer.php?mid=55). Scroll down to the heading "Covalent Bonding." OR (http://www.visionlearning.com/library/module_viewer.php?mid=55)

1. As opposed to _____ bonding in which a complete transfer of electrons occur, _____ bonding occurs when two or more elements _____ electrons. Covalent bonding occurs because the atoms in the compound have a similar tendency for electrons (generally to _____ electrons). This most commonly occurs when two _____ bond together. Because both of the nonmetals will want to _____ electrons, the elements involved will share electrons in an effort to _____ their valence shells.

2. How many valence electrons are in one atom of hydrogen?

3. How many valence electrons does hydrogen need to have a full first shell?

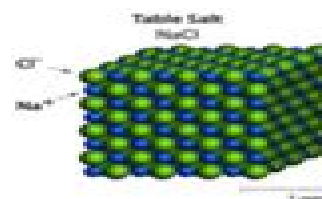
4. How does the hydrogen atom "pick up" another electron?

5. What compound does hydrogen form?

6. How do hydrogen atoms make a covalent bond?

7. Visit the simulation, [Covalent bonding between hydrogen atoms](http://www.visionlearning.com/library/flash_viewer.php?oid=1348&mid=55). OR (http://www.visionlearning.com/library/flash_viewer.php?oid=1348&mid=55) Describe or draw what you see.

8. Label the substances below as "ionic" or "covalent".



Answer Sheet To Overview Chemical Bonds

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Answer Sheet To Overview Chemical Bonds:

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Table of Contents Answer Sheet To Overview Chemical Bonds

1. Understanding the eBook Answer Sheet To Overview Chemical Bonds
 - The Rise of Digital Reading Answer Sheet To Overview Chemical Bonds
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet To Overview Chemical Bonds
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet To Overview Chemical Bonds
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet To Overview Chemical Bonds
 - Personalized Recommendations

- Answer Sheet To Overview Chemical Bonds User Reviews and Ratings
- Answer Sheet To Overview Chemical Bonds and Bestseller Lists
- 5. Accessing Answer Sheet To Overview Chemical Bonds Free and Paid eBooks
 - Answer Sheet To Overview Chemical Bonds Public Domain eBooks
 - Answer Sheet To Overview Chemical Bonds eBook Subscription Services
 - Answer Sheet To Overview Chemical Bonds Budget-Friendly Options
- 6. Navigating Answer Sheet To Overview Chemical Bonds eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Sheet To Overview Chemical Bonds Compatibility with Devices
 - Answer Sheet To Overview Chemical Bonds Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Sheet To Overview Chemical Bonds
 - Highlighting and Note-Taking Answer Sheet To Overview Chemical Bonds
 - Interactive Elements Answer Sheet To Overview Chemical Bonds
- 8. Staying Engaged with Answer Sheet To Overview Chemical Bonds
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Sheet To Overview Chemical Bonds
- 9. Balancing eBooks and Physical Books Answer Sheet To Overview Chemical Bonds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet To Overview Chemical Bonds
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Sheet To Overview Chemical Bonds
 - Setting Reading Goals Answer Sheet To Overview Chemical Bonds
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Sheet To Overview Chemical Bonds
 - Fact-Checking eBook Content of Answer Sheet To Overview Chemical Bonds

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the ion meta dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the

ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web stoichiometric principles can be used to determine the chemical equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab iron and copper sulfate sample data data and observations mass of copper ii sulfate pentahydrate 3 27 g of g mass of filter paper mass of filter paper and dry copper mass of dry copper 1 28 g