

1. Answer: A

**Explanation:** The value of  $\Delta H$  is negative. Therefore, the reaction is exothermic or heat-releasing.

2. Answer: A

**Explanation:** The entire reaction was multiplied by  $\frac{1}{2}$ . Therefore,  $\Delta H$  must also be halved.

3. Answer: C

**Explanation:** Option C is incorrect because  $-566 \text{ kJ/mol}$  is the  $\Delta H$  per mole of  $\text{O}_2$  and not for 2 moles. Even though options B and D are expressed in grams, they are still correct because 56 g of CO still corresponds to 2 moles of CO, and 32 g of  $\text{O}_2$  still corresponds to 1 mole of  $\text{O}_2$ , which is consistent with the expected factor.

4. Answer: A

**Explanation:**  $\Delta H = (98 \text{ g CO}_{(g)}) \left( \frac{1 \text{ mol CO}_{(g)}}{28 \text{ g CO}_{(g)}} \right) \left( \frac{-566 \text{ kJ}}{1 \text{ mol CO}_{(g)}} \right) = -990.5 \text{ kJ}$

5. Answer: B

**Explanation:**  $\Delta H = (1.5 \text{ mol CO}_{2(g)}) \left( \frac{-566 \text{ kJ}}{1 \text{ mol CO}_{2(g)}} \right) = -424.5 \text{ kJ}$

# Answer Key Unit 6 Thermochemistry

**Andrew Joseph Poss**



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### **Table of Contents Answer Key Unit 6 Thermochemistry**

1. Understanding the eBook Answer Key Unit 6 Thermochemistry
  - The Rise of Digital Reading Answer Key Unit 6 Thermochemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Unit 6 Thermochemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Key Unit 6 Thermochemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Unit 6 Thermochemistry
  - Personalized Recommendations
  - Answer Key Unit 6 Thermochemistry User Reviews and Ratings
  - Answer Key Unit 6 Thermochemistry and Bestseller Lists

5. Accessing Answer Key Unit 6 Thermochemistry Free and Paid eBooks
  - Answer Key Unit 6 Thermochemistry Public Domain eBooks
  - Answer Key Unit 6 Thermochemistry eBook Subscription Services
  - Answer Key Unit 6 Thermochemistry Budget-Friendly Options
6. Navigating Answer Key Unit 6 Thermochemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Answer Key Unit 6 Thermochemistry Compatibility with Devices
  - Answer Key Unit 6 Thermochemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answer Key Unit 6 Thermochemistry
  - Highlighting and Note-Taking Answer Key Unit 6 Thermochemistry
  - Interactive Elements Answer Key Unit 6 Thermochemistry
8. Staying Engaged with Answer Key Unit 6 Thermochemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Key Unit 6 Thermochemistry
9. Balancing eBooks and Physical Books Answer Key Unit 6 Thermochemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer Key Unit 6 Thermochemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Unit 6 Thermochemistry
  - Setting Reading Goals Answer Key Unit 6 Thermochemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Unit 6 Thermochemistry
  - Fact-Checking eBook Content of Answer Key Unit 6 Thermochemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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