

AP Chemistry Unit 3 Practice Test
Solids, Liquids, Gases and Solutions

Name _____
Date _____ Block _____

- A real gas would act most ideal at
(A) 1 atm and 273 K (D) 0.5 atm and 546 K
(B) 10 atm and 547 K (E) 0.5 atm and 273 K
(C) 10 atm and 273 K
- Zinc reacts with aqueous sulfuric acid to form hydrogen gas:
$$\text{Zn (s)} + \text{H}_2\text{SO}_4 \text{ (aq)} \rightarrow \text{ZnSO}_4 \text{ (aq)} + \text{H}_2 \text{ (g)}$$

In an experiment, 201 mL of wet H_2 is collected over water at 27°C and a barometric pressure of 733 torr. The vapor pressure of water at 27°C is 26.74 torr. The partial pressure of hydrogen in this experiment is _____ atm.

(A) 0.929 (D) 1.00
(B) 760 (E) 0.964
(C) 706
- A mixture of gases contains 1.5 moles of oxygen, 3.0 moles of nitrogen, and 0.5 mole of water vapor. If the total pressure is 700 mmHg, what is the partial pressure of the nitrogen gas?
(A) 70 mmHg (D) 350 mmHg
(B) 210 mmHg (E) 420 mmHg
(C) 280 mmHg
- Equal numbers of moles of $\text{CO}_2(\text{g})$, $\text{N}_2(\text{g})$, and $\text{NH}_3(\text{g})$ are placed in a sealed vessel at room temperature. If the vessel has a pinhole-size leak, which of the following will be true after some of the gas mixture has effused?
I. The mole fraction of CO_2 in the sample will increase.
II. The N_2 will effuse the fastest since it is the lightest.
III. All gases will effuse at the same rate since the temperature is held constant.
(A) I only (C) I and II (E) I, II, and III
(B) III only (D) II and III
- A sample of oxygen gas was found to effuse at a rate equal to two times that of an unknown gas. The molecular weight of the unknown gas is _____ g/mol.
(A) 128
(B) 8
(C) 16
(D) 64

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Richard H. Langley, John Moore



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