# AP Chemistry Chapter 1-4 Test Review

- What are the symbols for the elements: lead, silver, gold, potassium, mercury, iron, and sodium.
- What are the SI units for: mass, length, volume, energy, time?
- The density of aluminum is 2.7 g/cm<sup>3</sup>. A piece of aluminum that occupies a volume of 21.4 cm<sup>3</sup> would have a mass of?
- The number 1.00430 has \_\_\_\_\_ significant figures.
  A) 2 B) 3 C) 5 D) 6 E) 4
- 5. Identify the following as a chemical reaction or physical change:
  - A) Composting of grass
  - B) Evaporation of water
  - C) Rusting of a nail
  - D) Combustion of methane
  - E) Cutting paper
- 6. Identify which of the following are intensive physical properties and which are extensive:
  - A) Volume B) Mass C) Density D) Melting Point E) Color
- 7. What does the atomic number and atomic mass tell about the atom?
- 8. How do isotopes differ? Write the shorthand notation for two isotopes of an element.
- 9. Which elements are metals, which are metalloids and which are nonmetals?
- 10. What happens when a metal reacts with a nonmetal? What about a nonmetal and non metal?
- 11. What is the empirical formula for C<sub>4</sub>H<sub>8</sub>O<sub>2</sub>?
- 12. What are the typical oxidation numbers for each family?
- 13. Write the formula for Aluminum reacting with Oxygen.
- Name the following substances: HL CO<sub>2</sub>, MgCO<sub>3</sub>, H<sub>3</sub>PO<sub>4</sub>, SF<sub>4</sub>, MnO
- What are the major contributions of the following scientists to the atomic theory: Dalton, Thompson, Rutherford, Goldstein, Millikan, and Chadwick
- 16. How many electrons, neutrons and protons does 14C have?
- 17. What is the smallest particle of an atom?

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Neil D. Jespersen, Pamela Kerrigan

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