

AP Chemistry Chemical Equations Worksheet

Write the balanced chemical equation (including spectator ions underneath each reaction description, and answer the question.

(a) A solution of ammonia is added to a dilute solution of acetic acid. $\text{NH}_3\text{aq} + \text{CH}_3\text{COOH(aq)} \rightleftharpoons \text{NH}_4^+ + \text{CH}_3\text{COO}^-$	Identify the conjugate acid-base pairs in this reaction.
(b) Solutions of sodium hydroxide and acetic acid are mixed. $\text{NaCH}_3\text{COO} + \text{NaOH} \rightarrow \text{NaCH}_3\text{COONa} + \text{H}_2\text{O}$	Acetic acid is a weak acid. If equal volumes of equal molar solutions are mixed, will the solution pH be >7, 7, or <7? Explain.
(c) Hydrogen sulfide gas is bubbled through excess potassium hydroxide solution. $\text{H}_2\text{S} + 2\text{OH}^- \rightarrow \text{S}^{2-} + 2\text{H}_2\text{O}$	Write the successive ionization equations for H_2S.
(d) Solid barium oxide is added to distilled water.	Is the resulting solution acidic, basic, or neutral? Explain.
(e) Solid calcium oxide is exposed to a stream of carbon dioxide gas.	What type of reaction has occurred?
(f) Solid dinitrogen pentoxide is added to water.	Is the final solution acidic, basic, or neutral? Explain.
(g) Carbon disulfide vapor is burned in excess oxygen.	What is the oxidizing agent in this reaction?

Ap Chemistry Chemical Equations Worksheet Answers

Chris McMullen



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