

AP Chemistry - Chemical Kinetics

Name: _____ Date: _____ Period: _____

- After five half-life periods for a first order reaction, what fraction of reactant remains?
- If you plot $1/[\text{reactant}]$ versus time and observe a straight line, what is the order of the reaction? If $\ln[\text{reactant}]$ is plotted versus time, and a straight line of negative slope is observed, what is the order of the reaction?
- Indicate whether each of the following statements is true or false. Change the wording of each false statement to make it true.
 - In stepwise reactions the rate-determining step is the slow one.
 - It is possible to change the rate constant for a reaction by changing the temperature.
 - As a first order reaction proceeds at a constant temperature, the rate remains constant.
 - The rate constant for a reaction is independent of reactant concentrations.
- For each reaction below give the relative rates of formation of products and disappearance of reactants using coefficients and rate change symbols.
 - $2\text{O}_{3(g)} \rightarrow 3\text{O}_{2(g)}$
 - $2\text{HOF}_{(g)} \rightarrow 2\text{HF}_{(g)} + \text{O}_{2(g)}$
 - $\text{N}_{2(g)} + 3\text{H}_{2(g)} \rightarrow 2\text{NH}_{3(g)}$
- The reaction $\text{CO}_{(g)} + \text{NO}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{NO}_{(g)}$ is second order in NO_2 and zero order in CO at temperatures less than 500K.
 - Write the rate law for the reaction.
 - How will the reaction rate change if the NO_2 concentration is halved?
- For a reaction of the type $\text{A} + \text{B} \rightarrow \text{C}$, it is experimentally observed that doubling the concentration of B causes the reaction rate to be increased fourfold, but doubling the concentration of A has no effect on the rate. Determine the initial rate equation.
- A study of the reaction $2\text{A} + \text{B} \rightarrow \text{C} + \text{D}$ gave the following results:

Expt	Init Con (M)		Init Rate (M/s)
	[A] ₀	[B] ₀	
1.	0.10	0.05	6.0×10^{-3}
2.	0.20	0.05	1.2×10^{-2}
3.	0.30	0.05	1.8×10^{-2}
4.	0.20	0.15	1.1×10^{-1}

- What is the rate law for this reaction?
 - What is the overall order of the reaction?
 - Could this reaction occur in a single step?
- The decomposition of N_2O_5 (to give NO_2 and O_3) follows the rate law "rate = $k[\text{N}_2\text{O}_5]^n$ " where k is 5.0×10^{-4} per second at a particular temperature.
 - What is the half-life of this reaction?
 - How long does it take for the N_2O_5 concentration to drop to one tenth of its original value?

Ap Chemistry Chemical Kinetics Worksheet Answers

Albert S. Tarendash, Frederick J. Rowe



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In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Ap Chemistry Chemical Kinetics Worksheet Answers**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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