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web stant is the foundation of thermochemistry it allows us to describe heat flow in terms of the change in a state function
enthalpy change is path independent note with regard to choice b that enthalpy change equals internal energy change only if
the volume of the system remains constant and thus no thermodynamic work is performed 5 b

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web conceptual answers in order for the reaction to occur spontaneously ΔG for the reaction must be less than zero in this
case ΔS must be positive and the $T\Delta S$ term outweighs the positive value of ΔH

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increases to 58.6 °C from an initial temperature of 22.1 °C these are homework exercises to accompany the textmap created for
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web 1 answer a explanation the value of ΔH is negative therefore the reaction is exothermic or heat releasing 2 answer a 1
explanation the entire reaction was multiplied by 2 therefore ΔH must also be halved 3 answer c explanation option c is
incorrect because 566 kJ/mol is the ΔH per mole of O_2 and not for 2 moles

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web thermochemical equations equations that represent both mass and enthalpy changes $H_2O(s) \rightarrow H_2O(l)$ 6.01 kJ/mol this is
an endothermic process it requires 6.01 kJ to melt one mole of ice $H_2O(s)$ the enthalpy value will change if the number of
moles varies from the 1:1 reaction stoichiometry this is an exothermic process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy e of a system in any process is equal to the heat q added to the system plus the work w done on the system by its surroundings e

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