

Question 1

1(a)(iii) Difference: With powdered CaCO_3 the graph has a steeper slope at the start of the reaction than with marble chips because the rate of reaction is faster due to increased surface area of the solid.
Similarity: The total volume of CO_2 produced is the same in both cases since the same mass (and therefore number of moles) of CaCO_3 was used in both experiments.



1(a)(v) Moles of CaCO_3 used = $(1.8/100)$ mol. = 0.018 mol.
From the balanced equation for the reaction, 1 mole CaCO_3 produces 1 mol CO_2 .
Hence volume of gas produced at R.T.P. = (0.018×24000) cm^3 = 432 cm^3 .

1(a)(vi) There were impurities in the CaCO_3 used.

1(a)(vii) Concentration of the acid used and temperature at which the reaction was carried out.

1(b)(i) $\text{Fe}^{3+}_{(\text{aq})}$ present.

1(b)(ii) Magnesium ribbon decreases in size/dissolves.
Solution becomes pale green.



1(b)(iv) SO_4^{2-} , SO_3^{2-} and CO_3^{2-} absent.

1(b)(v) White precipitate, which is soluble in aqueous ammonia, is formed.

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