

Experiment 10: Electrochemistry with a Standard Hydrogen Electrode from: *Basic Electrochemistry*, Experiment 10, pp. 45-50

The purpose of this experiment is to determine standard potentials for several other electrodes. In the experiment, a standard hydrogen electrode (SHE) is used as a reference electrode. The SHE is a half-cell consisting of a platinum electrode immersed in a 1 M H^+ solution, with H_2 gas at 1 atm pressure. The potential of the SHE is defined as 0 V. Other electrodes are connected to the SHE, and their standard potentials are determined relative to the SHE. The standard potential of a half-cell is a measure of its tendency to be oxidized or reduced. The standard potential of a half-cell is determined by measuring the potential difference between the half-cell and the SHE.

In this experiment, several half-cells are connected to the SHE, and their standard potentials are determined. The half-cells are: (1) a zinc electrode in a 1 M Zn^{2+} solution, (2) a copper electrode in a 1 M Cu^{2+} solution, (3) a silver electrode in a 1 M Ag^+ solution, and (4) a cadmium electrode in a 1 M Cd^{2+} solution.

Objectives:

1. To determine the standard potential of the zinc electrode.
2. To determine the standard potential of the copper electrode.
3. To determine the standard potential of the silver electrode.
4. To determine the standard potential of the cadmium electrode.
5. To determine the standard potential of the SHE.

Materials:

- Standard hydrogen electrode (SHE)
- Zinc electrode
- Copper electrode
- Silver electrode
- Cadmium electrode
- 1 M Zn^{2+} solution
- 1 M Cu^{2+} solution
- 1 M Ag^+ solution
- 1 M Cd^{2+} solution
- 1 M H^+ solution

Procedure:

1. Connect the SHE to the zinc electrode.
2. Measure the potential difference between the SHE and the zinc electrode.
3. Repeat the procedure for the other electrodes.

Results:

1. The standard potential of the zinc electrode is -0.76 V.
2. The standard potential of the copper electrode is +0.34 V.

Discussion:

1. The standard potential of the zinc electrode is negative, indicating that it is more likely to be oxidized than the SHE.
2. The standard potential of the copper electrode is positive, indicating that it is more likely to be reduced than the SHE.

Answers For Lab 6b Heath Chemistry

RD Boyd



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