

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation: $\underline{\quad 2 \quad} \text{H}_2\text{S}_{(g)} + \underline{\quad 3 \quad} \text{O}_{2(g)} \rightarrow \underline{\quad 2 \quad} \text{SO}_{2(g)} + \underline{\quad 2 \quad} \text{H}_2\text{O}_{(g)}$

Before $\underline{\quad 2 \quad}$ $\underline{\text{ex}}$ $\underline{\quad 0 \quad}$ $\underline{\quad 0 \quad}$ 3.0 mol O₂

Change $\underline{\quad -2 \quad}$ $\underline{\quad -3 \quad}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

After $\underline{\quad 0 \quad}$ $\underline{\text{ex}}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

- Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$

Before 0.5 $\underline{\text{ex}}$ 0 0 1.50mol HCL

Change -0.5 -1.5 +0.5 +0.5

After 0 $\underline{\text{ex}}$ 0.5 0.75

- Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before $\underline{\text{ex}}$ 1.25 0 2.50 mole

Change -3.75 -1.25 +2.5

After $\underline{\text{ex}}$ 0 2.5

Answers Unit 8 1 Mole Relationships Answers

Peter P. Berlow



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Solving Problems in Chemistry Smith,1993-06 **Report of Investigations** ,1962 *Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed* Peter P. Berlow,1978 **Introduction to Chemistry** Thomas R. Dickson,1983 Teaches chemistry by offering a dynamic provocative relevant view of the topic its importance to society our daily lives Three themes are stressed throughout the text developing chemical thinking a chemical vision learning problem solving methods utilizing group work discussion activities These themes involve engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking shows that chemistry involves interesting relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment society as well as to specific tools of direct use to students Educart CBSE Question Bank Class 11 Chemistry 2024-25 (For 2025 Board Exams) Educart,2024-06-17 What You Get Time Management ChartsSelf evaluation ChartCompetency based Q sMarking Scheme Charts Educart Class 11 Chemistry Strictly based on the latest CBSE Curriculum released on March 31st 2023Related NCERT theory with diagrams flowcharts bullet points and tablesImportant and Caution Points give to really work on common mistakes made during the examLots of solved questions with Detailed Explanations for all questionsIncludes Case based Examples and Numerical based Questions as per the new pattern changeExtra practice questions from various CBSE sources such as DIKSHA platform and NCERT exemplars Why choose this book You can find the simplified complete with diagrams flowcharts bullet points and tablesBased on the revised CBSE pattern for competency based questionsEvaluate your performance with the self evaluation charts Cambridge International AS and A Level Chemistry Revision Guide Judith Potter,Peter Cann,2015-10-29 A revision guide tailored to the AS and A Level Chemistry syllabus 9701 for first examination in 2016 This Revision Guide offers support for students as they prepare for their AS and A Level Chemistry 9701 exams Containing up to date material that matches the syllabus for examination from 2016 and packed full of guidance such as Worked Examples Tips and Progress Check questions throughout to help students to hone their revision and exam technique and avoid common mistakes These features have been specifically designed to help students apply their knowledge in exams Written in a clear and straightforward tone this Revision Guide is perfect for international learners **Molten-Salt Reactor Program Progress Report for Period from ...** Oak Ridge National Laboratory,1960-08 **Chemistry Class - XII Model Paper Chapter wise Question Answer With Marking Scheme 2023- SBPD Publications** SBPD Editorial Board,2022-10-14 Content 1 Solid State 2 Solution 3 Electrochemistry 4 Chemical Kinetics 5 Surface Chemistry 6 General Principles and Processes of Isolation of Element 7 P Block elements 8 d and f Block Elements 9 Coordination Compounds 10 Haloalkanes and Haloarenes 11 Alcohols Phenols and Ethers 12 Aldehydes Ketones and Carboxylic Acid 13 Organic

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