AP Chem Unit 2 a

bonds

-based on electronegativity increpad.

*electronegativity= measure of how well atoms altract shared electroms.

x - xX - Ypolem nompolar (same element) (sist element) * polar covalent poind arrown with greater electromore negativity will e sec trromoga hwe have parrial * greater diff. megative charge in etectroreq

-interestitial alloy: significant difference in atomic radius > smaller atoms fill in gaps

attivity evening

Pohous

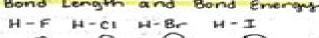


substitutional alloy-2 elements in which the atomic radius I'D SIMMILANT



- metal + nonmetal = ronic
- nonmetal +nonmetal = covalent
- good conductor of electricity: metallic bonding
- solid doesn't conduct but as liquid does = ionic bonding
- solid that doesn't conduct flow melting point *covalent molecular
- solid that doesn't conduct ! high meiting point covalent network *metallic bonding: sea of mobile valence electrons @

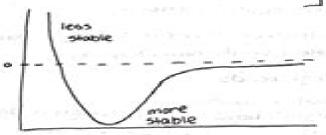
and Bond Energy



- Shortest lerond. - hand to bereatt.
- C CC = C- goding to burneou/se... -longest
- -lengest bornd - 605% 10 itamedak.

out bond energy increases CFC it becomes honder - hand to to break and bond bureout. langth is sharten - sever hear. Done

4 good conductors



lewis dot diagrams

CH2Cl2 = 20 valence e -

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