

1999 Advanced Placement Chemistry Exam

Section I: Multiple Choice

Time — 1 hour and 30 minutes

NO CALCULATORS MAY BE USED WITH SECTION I.

Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified.

Throughout the test the following symbols have the definitions specified unless otherwise noted.

T = temperature
 P = pressure
 V = volume
 S = entropy
 H = enthalpy
 G = free energy
 R = molar gas constant
 n = number of moles
 M = molar

m = molal
 L, mL = liter(s), milliliter(s)
 g = gram(s)
 nm = nanometer(s)
 atm = atmosphere(s)
 J, kJ = joule(s), kilojoule(s)
 V = volt(s)
 mol = mole(s)

Part A

Directions: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then fill in the corresponding oval on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–4 refer to the following types of energy.

- (A) Activation energy
- (B) Free energy
- (C) Ionization energy
- (D) Kinetic energy
- (E) Lattice energy

1. The energy required to convert a ground-state atom in the gas phase to a gaseous positive ion
2. The energy change that occurs in the conversion of an ionic solid to widely separated gaseous ions
3. The energy in a chemical or physical change that is available to do useful work
4. The energy required to form the transition state in a chemical reaction

Questions 5–8 refer to atoms for which the occupied atomic orbitals are shown below.

- (A) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (B) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (C) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (D) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (E) $[Ar] 4s \underline{\hspace{1cm}} 3d \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$

5. Represents an atom that is chemically unreactive
6. Represents an atom in an excited state
7. Represents an atom that has four valence electrons
8. Represents an atom of a transition metal

Questions 9–12 refer to aqueous solutions containing 1:1 mole ratios of the following pairs of substances. Assume all concentrations are 1 M .

Ap Chemistry 2007 Answers

David Wilson



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Table of Contents Ap Chemistry 2007 Answers

1. Understanding the eBook Ap Chemistry 2007 Answers
 - The Rise of Digital Reading Ap Chemistry 2007 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry 2007 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Chemistry 2007 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry 2007 Answers
 - Personalized Recommendations
 - Ap Chemistry 2007 Answers User Reviews and Ratings
 - Ap Chemistry 2007 Answers and Bestseller Lists

5. Accessing Ap Chemistry 2007 Answers Free and Paid eBooks
 - Ap Chemistry 2007 Answers Public Domain eBooks
 - Ap Chemistry 2007 Answers eBook Subscription Services
 - Ap Chemistry 2007 Answers Budget-Friendly Options
6. Navigating Ap Chemistry 2007 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ap Chemistry 2007 Answers Compatibility with Devices
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ap Chemistry 2007 Answers
 - Highlighting and Note-Taking Ap Chemistry 2007 Answers
 - Interactive Elements Ap Chemistry 2007 Answers
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ap Chemistry 2007 Answers
9. Balancing eBooks and Physical Books Ap Chemistry 2007 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Chemistry 2007 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ap Chemistry 2007 Answers
 - Setting Reading Goals Ap Chemistry 2007 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chemistry 2007 Answers
 - Fact-Checking eBook Content of Ap Chemistry 2007 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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