

Unit 1 Atomic Structure & Properties

KEY terms

Atomic Mass: The mass of an element
Molar Mass: The mass of a mole worth of an element
Mass Percent: The percent in mass of something of the overall
 Avogadro's # $6.02 \cdot 10^{23}$
Isotopes: Different types of an element - different number of neutrons
Empirical formula: reduced version of a formula

Atomic Mass

Where does it come from?

Each element has at least one isotope, and each isotope has a different weight, to find the atomic mass, scientists averaged the weights of each isotope of an element based on how much of each isotope exists in nature.

But what is an Isotope?

Isotopes are different forms of an element and each element has at least one.

An element is determined based on its number of protons. Electrons determine how reactive an element is.

The number of neutrons which isotope it is.

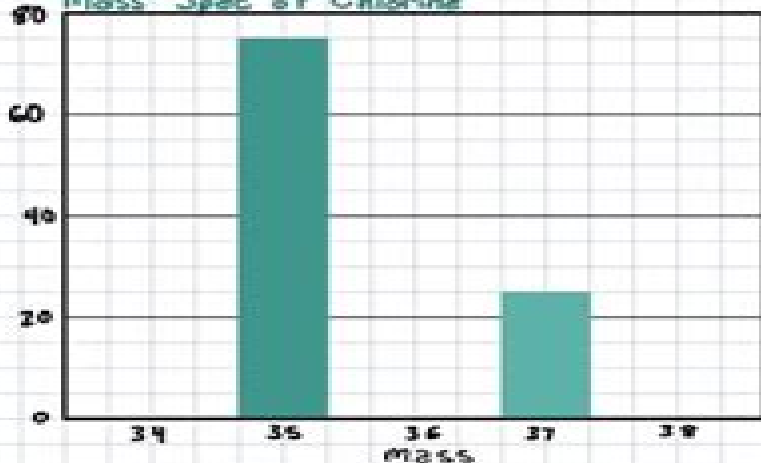
Mass Spectroscopy on the AP Exam

Estimate atomic mass from a mass spectrum graph
 (could be a fake element)

Identifying an element from a mass spectrum graph

Identifying an isotope from a mass spectrum graph

Mass Spec of Chlorine



Approx. Avg Atomic Mass

$\approx 75\%$ of the isotopes found have a mass of 35 grams

$\approx 25\%$ of the isotopes found have a mass of 37 grams

$$\frac{75\%}{100\%} = 0.75 \quad \frac{25\%}{100\%} = 0.25$$

$$(35 \cdot 0.75) + (37 \cdot 0.25)$$

$$26.25 + 9.25$$

$$35.5$$

sig figs (36g)

Estimate the atomic mass because most of the mass is 35 grams the element needs to be about 35 grams

18	19	20
Ar	K	Ca
Argon	Potassium	Calcium

Potassium makes the most sense as the closest to 39 grams

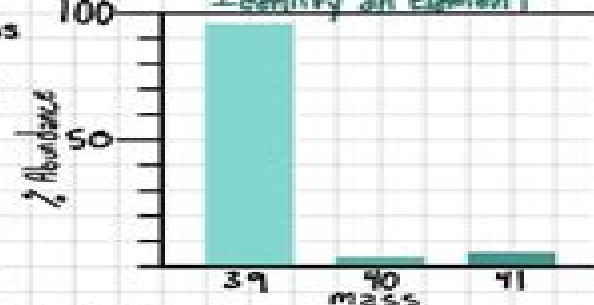
PERIODIC TABLE OF THE ELEMENTS

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99	100
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- Alkali metals
- Alkaline earth metals
- Transition metals
- Post-transition metals
- Nonmetals
- Reactive nonmetals
- Halogen gases
- Lanthanides
- Actinides
- Unlabeled properties

Identify an Element



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Judd E. Hollander



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