

A P Chemistry

Chapter 13 Pre-Test

- ____ 1. Indicate the mass action expression for the following reaction:



- A) $\frac{[3W][V]}{[2X][Y]}$ B) $\frac{[X]^2[Y]}{[W]^3[V]}$ C) $\frac{[W]^3[V]}{[X]^2[Y]}$ D) $[X]^2[Y][W]^3[V]$

- ____ 2. The value of the equilibrium constant, K , is dependent on
- I. The temperature of the system.
 - II. The nature of the reactants and products.
 - III. The concentration of the reactants.
 - IV. The concentration of the products.
- A) I, II
 B) II, III
 C) III, IV
 D) It is dependent on three of the above choices.
 E) It is not dependent on any of the above choices.

- ____ 3. Apply the law of mass action to determine the equilibrium expression for



- A) $[NO_2Cl]^2/[NO_2]^2[Cl_2]$ D) $2[NO_2Cl]/2[NO_2][Cl_2]$
 B) $2[NO_2][Cl_2]/2[NO_2Cl]$ E) $[NO_2Cl]^2/[NO_2]^2[Cl_2]$
 C) $[NO_2]^2[Cl_2]/[NO_2Cl]^2$

Use the following to answer questions 4-5:

Consider the chemical system $CO + Cl_2 \rightleftharpoons COCl_2$;
 $K = 4.6 \times 10^9 \text{ L/mol}$.

- ____ 4. If the concentration of the product were to double, what would happen to the equilibrium constant?
- A) It would double its value.
 B) It would quadruple its value.
 C) It would not change its value.
 D) It would depend on the initial conditions of the product.
 E) It would become half its current value.
- ____ 5. Determine the equilibrium constant for the system $N_2O_4 \rightleftharpoons 2NO_2$ at 25°C. The concentrations are shown here: $[N_2O_4] = 4.27 \times 10^{-2} \text{ M}$, $[NO_2] = 1.41 \times 10^{-2} \text{ M}$
- A) 0.66 B) 3.0 C) 0.0047 D) 0.33 E) 0.05
- ____ 6. Consider the reaction $H_2 + I_2 \rightleftharpoons 2HI$ whose $K = 54.8$ at 425°C. If an equimolar mixture of reactants gives the concentration of the product to be 0.50 M at equilibrium, determine the concentration of the hydrogen.
- A) $1.6 \times 10^{-4} \text{ M}$ B) $4.6 \times 10^{-3} \text{ M}$ C) $9.6 \times 10^{-2} \text{ M}$ D) $6.8 \times 10^{-2} \text{ M}$ E) $1.2 \times 10^{-3} \text{ M}$

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John Moore, Richard H. Langley



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