AP Chemistry: Chapter 14 Practice Test

10. Given the following acids and Ka values:

HClO₄ HOAe

HCN

HF

 1×10^7 1.76×10^{-6} 4.93×10^{-10} 3.53×10^{-4} which shows the conjugate bases listed by increasing strength?

[A] ClO₄-, F-, OAc-, CN-

[B] CN-, F-, OAc-, ClO₄-

[C] CN-, OAe-, F-, ClO₄-

[D] CN-, ClO₄-, F-, OAc-

[E] ClO₄-, OAc-, CN-, F-

Given that the K₈ for HOCl is 3.5 × 10⁻⁸, calculate the K value for the reaction of HOCl with OH...

[A] 3.5×10^{-22}

[B] 3.5×10⁶

 1012.9×10^{-7}

(D) 3.5×10^{-6}

[E] none of these

Acetic acid (HC₂H₃O₂) is a weak acid (K_a = 1.8 × 10−5). Calculate the pH of a 17.6 M HC₂H₃O₂ solution.

[A] 4.3

[B] 6.4

PC1 1.7

[D] 0.97

[E] 7.4

What concentration of acetic acid (K_a = 1.80 × 10⁻⁵) has the same pH as that of 5.00 × 10⁻⁵ M HCl?

 $[A] 5.00 \times 10^{-3} M$

[B] 3.60 × 10⁻³ M

[C] 1.39 M.

[D]: 5.00 M

(E) none of these

Consider the reaction HOCl + F - HF + OCl

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John T. Moore, Richard H. Langley

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