

AP* Chemistry ELECTROCHEMISTRY

Terms to Know:

- Electrochemistry**—the study of the interchange of chemical and electrical energy
OIL RIG – oxidation is loss, reduction is gain (of electrons)
Oxidation – the loss of electrons, increase in charge
Reduction – the gain of electrons, reduction of charge
Oxidation number – the assigned charge on an atom
Oxidizing agent (OA) – the species that is reduced and thus *causes* oxidation
Reducing agent (RA) – the species that is oxidized and thus *causes* reduction

ELECTROCHEMISTRY INVOLVES TWO MAIN TYPES OF PROCESSES:

- Galvanic (voltaic) cells** – which are spontaneous chemical reactions (battery)
- Electrolytic cells** – which are non-spontaneous and require external e^- source (DC power source)
- BOTH** of these fit into the category entitled **Electrochemical cells**

GALVANIC CELLS

Parts of the voltaic or galvanic cell:

- o Anode – the electrode where oxidation occurs. After a period of time, the anode may appear to become smaller as it falls into solution.
- o Cathode – the anode where reduction occurs. After a period of time it may appear larger, due to ions from solution plating onto it.
- o inert electrodes—used when a gas is involved OR ion to ion involved such as Fe^{3+} being reduced to Fe^{2+} rather than Fe^0 . Made of Pt or graphite.
- o Salt bridge – a device used to maintain electrical neutrality in a galvanic cell. This may be filled with agar which contains a neutral salt or it may be replaced with a porous cup.
- o Electron flow -- always from anode to cathode. (through the wire)
- o Standard cell notation (line notation) - anode/solution// cathode solution/ cathode Ex $Zn/Zn^{2+} (1.0 M) // Cu^{2+} (1.0 M) / Cu$
- o Voltmeter - measures the cell potential (ϵ_{cell}). Usually is measured in volts.

Balance this re-dox reaction:



RED:

OA:

OX:

RA:

Overall rxn:



Ap Chemistry Electrochemistry Answers

J Elliott



Ap Chemistry Electrochemistry Answers:

The book delves into Ap Chemistry Electrochemistry Answers. Ap Chemistry Electrochemistry Answers is an essential topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Ap Chemistry Electrochemistry Answers, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Ap Chemistry Electrochemistry Answers
- Chapter 2: Essential Elements of Ap Chemistry Electrochemistry Answers
- Chapter 3: Ap Chemistry Electrochemistry Answers in Everyday Life
- Chapter 4: Ap Chemistry Electrochemistry Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Ap Chemistry Electrochemistry Answers. This chapter will explore what Ap Chemistry Electrochemistry Answers is, why Ap Chemistry Electrochemistry Answers is vital, and how to effectively learn about Ap Chemistry Electrochemistry Answers.
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6. In chapter 5, the author will draw a conclusion about Ap Chemistry Electrochemistry Answers. The final chapter will summarize the key points that have been discussed throughout the book.

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