

Unit 1 Review

1. What is the difference between the three types of representative democracies: participatory democracy, pluralist democracy, and elite democracy?
 - Participatory – the people directly have the power to make policy
 - Pluralist – different groups compete/have power to make policy
 - Elite – a small group of people who have more power, make policy
2. Give examples of how participatory democracy is presented in American government.
State/local levels of government – initiative, referendum, & recall
 - Initiative – the people initiate/begin a law themselves
 - Referendum – the people vote directly on a law
 - Recall – the people can vote an elected official out of office before their term ends
3. Give examples of how pluralist democracy is presented in American government.
 - Interest groups
 - Political parties
4. Give examples of how elite democracy is presented in American government.
 - Old, white, rich, Protestant, men = Congress
 - Electoral College (electors choose the President)
5. Define popular sovereignty. How is this reflected in American government?
 - Consent of the governed/Government is created and given its power by the people
 - We choose our representatives (Congress members)
6. Define limited government. How is this reflected in American government?
 - A government that has limited powers (Constitution will limit the powers of the government)
 - Separation of powers, checks and balances, federalism, Bill of Rights
7. Define social contract. How is this reflected in American government?
 - Agreement between the people and government (people will agree to give up some freedoms/rights in order to be protected by the government)
 - Bill of Rights is not absolute (ex: can't yell fire in crowded movie theater), PATRIOT ACT after 9/11
8. Define republic. How is this reflected in American government?
 - Political system where people choose representatives to represent them
 - Electing our Congress members
9. What Enlightenment philosopher came up with the idea of natural rights? According to this person

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web nov 23 2017 maharashtra x physics suppose the masses of calorimeter th water in it and the hot object made up of copper which is put in the calorimeter are the same the initial temperature of the calorimeter and water is 30 c and that of hot object is 60 c the specific heat of copper and water are 0 09 cal gm c and 1 cal gm c respectively

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web oct 28 2019 solution use this equation $q = \text{specific heat} \times m \times \Delta t$ where q is heat flow m is mass in grams and Δt is the temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ J/g}^\circ\text{C} \times 110 \text{ g} \times 26.6^\circ\text{C} = 12500 \text{ J}$ $q_{\text{water}} = 550 \text{ J}$ $\Delta h = q_{\text{water}} = 550 \text{ J}$

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web calorimetry problems calorimetry problems $q_{\text{sur}} = m \times c \times \Delta t$ q heat m mass $q_{\text{rxn}} = q_{\text{sur}} + t_f + t_i$ name per date c specific heat for water 4.184 J/g°C 1 what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased by 8.1°C when 207 J of heat is added 2

7 3 heats of reactions and calorimetry chemistry libretexts - Dec 30 2022

web jul 12 2023 this concept lies at the heart of all calorimetry problems and calculations because the heat released or absorbed at constant pressure is equal to ΔH the relationship between heat and ΔH_{rxn} is $\Delta H_{\text{rxn}} = q_{\text{calorimeter}} / m \times c \times \Delta t$ 7 3 14 7 3 14 $\Delta H_{\text{rxn}} = q_{\text{rxn}} / m \times c \times \Delta t$

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web use the information in this chart to answer the following questions 1 if i have 125 grams of water at a temperature of 25°C and heat it to 145°C what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the process in 1 3 i have 18.7 grams of an unknown liquid chemical at a temperature of 35.0°C

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web sample problem calorimetry and enthalpy changes in an experiment 25.0 ml of 1.00 M HCl at 25.0°C is added to 25.0 ml of 1.00 M NaOH at 25.0°C in a foam cup calorimeter a reaction occurs and the temperature rises to 32.0°C calculate the enthalpy change in kJ for this reaction assume the densities of the solutions are 1.00 g/ml and that

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web before we practice calorimetry problems involving chemical reactions consider a simple example that illustrates the core idea behind calorimetry suppose we initially have a high temperature substance such as a hot piece of metal m and a low temperature substance such as cool water w answer the initial temperature of the copper

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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web calorimetry worksheet 1 if 0 315 moles of hexane C_6H_{14} is combusted in a bomb calorimeter containing 5 65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 55 4 $^{\circ}\text{C}$ the specific heat capacity of water is 4 184 J $\text{g}^{-1} \text{ } ^{\circ}\text{C}^{-1}$ $m_{\text{H}_2\text{O}} = 5 650 \text{ g}$ $h_{\text{H}_2\text{O}} = 4 184 \text{ J } ^{\circ}\text{C}^{-1} \text{ g}^{-1}$ $h_{\text{hexane}} = 1310 \text{ kJ}$

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web 6 a 218 g sample of steam at 121 $^{\circ}\text{C}$ is cooled to ice at 14 $^{\circ}\text{C}$ find the change in heat content of the system 7 in going from ice at 34 $^{\circ}\text{C}$ to steam at 138 $^{\circ}\text{C}$ a sample of water absorbs $1 41 \times 10^5 \text{ J}$ find the mass of the sample answers 1 2 68 $\times 10^4 \text{ J}$ 2 6 23 $\times 10^4 \text{ J}$ 3 1 71 $\times 10^4 \text{ J}$ 4 1 11 $\times 10^5 \text{ J}$ 5 2 11 $\times 10^5 \text{ J}$ 6

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