Chemistry

(www.tiwariacademy.com) (Chapter 1)(Solutions) (Intext Questions)

(Class 12)

Question 1.1:

Calculate the mass percentage of benzene (C₆H₆) and carbon tetrachloride (CCl₆) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride.

Answer 1.1:

Mass percentage of $C_6H_6 = \frac{Mass \ of \ C_6H_6}{Total \ mass \ of \ the \ solutions} \times 100 \%$

$$= \frac{Mass\ of\ C_6 H_6}{Mass\ of\ C_6 H_6 + Mass\ of\ CCl_4} \times 100\ \% = \frac{22}{22 + 122} \times 100\ \% = 15.28\ \%$$

Mass percentage of $CCl_4 = \frac{Mass\ of\ CCl_4}{Total\ mass\ of\ the\ solutions} \times 100\ \%$

$$= \frac{Mass\ of\ CCl_4}{Mass\ of\ C_6H_6 + Mass\ of\ CCl_4} \times 100\ \% = \frac{122}{22 + 122} \times 100\ \% = 84.72\ \%$$

Alternatively, Mass percentage of CCl₄ = (100 - 15.28) % = 84.72%

Question 1.2:

Calculate the mole fraction of benzene in solution containing 30% by mass in carbon tetrachloride.

Answer 1.2:

Let the total mass of the solution be 100 g and the mass of benzene be 30 g.

. Mass of carbon tetrachloride =
$$(100 - 30)$$
 g = 70 g

Molar mass of benzene (CsH₆) = $(6 \times 12 + 6 \times 1)$ g mol⁻¹ = 78 g mol⁻¹

. Number of moles of
$$C_6H_6 = \frac{30}{78}$$
 mol = 0.3846 mol

Molar mass of carbon tetrachloride (CCI₄) = $1 \times 12 + 4 \times 355$ = 154 g mol^{-1}

.Number of moles of
$$CCl_4 = \frac{70}{154}$$
 mol = 0.4545 mol

Thus, the mole fraction of CoHo is given as:

Number of moles of C₆H₆ + Number of moles of CCl₄
0.3846

- 0.3846 + 0.4545
- = 0.458

Question 1.3:

Calculate the molarity of each of the following solutions:

Answer 1.3:

Molarity is given by: $Molarity = \frac{Moles \ of \ solution \ in \ litre}{Volume \ of \ solution \ in \ litre}$

- (a) Molar mass of Co (NO₃)₂.6H₂O = $59 + 2(14 + 3 \times 16) + 6 \times 18 = 291 \text{ g mol}^{-1}$
 - .. Moles of Co (NO₃)₂.6H₂O = $\frac{30}{291}$ mol = 0.103 mol

Therefore, molarity =
$$\frac{0.103 \text{ mol}}{4.3 \text{ L}} = 0.023 \text{ M}$$

(b) Number of moles present in 1000 mL of 0.5 M H₂SO₄ = 0.5 mol

. Number of moles present in 30 mL of 0.5 M
$$H_2SO_4 = \frac{0.5\times30}{1000}$$
 mol = 0.015 mol

Therefore, molarity =
$$\frac{0.015 \text{ mol}}{0.5 \text{ L}} = 0.03 \text{ M}$$

8 Chemistry Sol Answers

Kohli Nitin

8 Chemistry Sol Answers:

Solutions Manual to Accompany Elements of Physical Chemistry David Smith, 2013-05-30 The Solutions Manual to accompany Elements of Physical Chemistry 6th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these Solutions Manual to Accompany Physical Chemistry for the Life Sciences C. A. Trapp, M. P. Cady, 2011 This questions solutions manual contains fully worked solutions to all end of chapter discussion questions and exercises featured in Physical Chemistry for the Life Sciences Longman Science Chemistry 9 Kohli Nitin, 2008-09 **Excel With Objective Chemistry** For Iit-Screening Prof. S. K. Khanna, Dr. N. K. Sharma, Dr. B. Kapila, 2006 Chapter-wise NCERT + Exemplar + Practice Questions with Solutions for CBSE Chemistry Class 11 - 2nd Edition Disha Experts, 2017-08-29 The book Chapter wise NCERT Exemplar Practice Questions with Solutions for CBSE Class 11 Chemistry has been divided into 3 parts Part A provides detailed solutions Question by Question of all the questions exercises provided in the NCERT Textbook Part B provides solutions to the questions in the NCERT Exemplar book Part C provides selected Practice Questions useful for the Class 11 examination along with detailed solutions The solutions have been designed in such a manner Step by Step that it would bring 100% Concept Clarity for the student Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction The Best Books William Swan Sonnenschein, 1895 JAM IIT Chemistry Question Bank MSC Entrance Mocktime Publication, The IIT JAM Chemistry code CY guestion bank includes topic wise drills Inorganic Organic Physical Chemistry Analytical Environmental Chemistry previous year JAM papers in PDF with solutions books with aligned notes and model test series The exam pattern for JAM Chemistry consists of 60 questions 100 marks across Sections A B C MCQs MSQs NATs consistent with IIT JAM structure The question bank aids entrants by exposing them to high frequency topics helping develop speed and accuracy and offering revision material and full length tests that mirror actual exam conditions Ideal for MSc entrance The Chemical News and Journal of Physical Science, 1893 preparation in top institutes A Reader's Guide to the Choice of the Best Available Books (about 50,000) in Every Department of Science, Art & Literature, with the Dates of the First & Last Editions, & the Price, Size & Publisher's Name of Each Book William Swan Sonnenschein, 1901 A Reader's Guide to Contemporary Literature William Swan Sonnenschein, 1895 **33 Years Chapterwise Solutions NEET**

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