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IGCSE A* MODEL ANSWER

0620 CHEMISTRY

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Paper 42



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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.



Table 3.1 shows some bond energies.

Table 3.1

bond	H-H	N-H
bond energy in kJ/mol	436	391

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an O-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.
 $(40H) + 1(N_2)$
 $4(436) + 1(945)$

3350 kJ
- Calculate the energy released when bonds form in the products.
 $2(6N-H) + 4(2O-H) = -90$
 $2(2350) + 4(463) = -90$

2340 kJ
- Calculate the energy of an O-H bond.
 $6 \times (N-H) = 2340$
 $N-H = \frac{2340}{6}$

390 kJ/mol

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