- A piece of aluminum foil 1.00 cm square and 0.550 mm thick is allowed to react with bromine to form aluminum bromide.
 - a.) How many moles of aluminum were used? (density Al = 2.699 g/cm²)

b.) How many grams of aluminum bromide form, assuming the aluminum reacts completely? What type of reaction is this? ZA1 +3BC, -> ZA1BG

Determine the percent composition of H and C for benzene

$$G(13.01) + G(1.01) = 48.12$$

$$C: \frac{73.06}{78.13} \times 100 = 93.70 \quad H = 100-93.70 = 8.70$$
Write a balanced combustion reaction for benzene

Write a balanced combustion reaction for benzene.

c. Assuming you have 364g of benzene, how many moles of O2 will be required to completely combust this amount of benzene?

d. How many grams of CO₂ will this produce? How many grams of H₂O?

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John Olmsted, Greg Williams, Robert C. Burk

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